

# FORMULAS AND OXIDATION NUMBERS

A chemical formula is a combination of symbols and numerical subscripts that represents the composition of a compound. The symbols indicate which elements are present and the numerical subscripts indicate the relative proportion of each element in the compound. These proportions can be predicted using the oxidation numbers of the elements and the charges of polyatomic ions. When atoms acquire a charge they are called ions. Ions consisting of more than one atom are polyatomic ions. Its oxidation number represents the apparent charge on an atom.

It is important that all scientists use the same system for writing chemical formulas. This helps to ensure clear and consistent transmission of information. Therefore, the following rules should be used for writing chemical formulas.

1. In a neutral compound the sum of the oxidation numbers of the elements and the charges on polyatomic ions in that compound must equal zero.
2. One positive (+) charge will neutralize one negative (-) charge.
3. Atoms with positive oxidation numbers or ions with positive charges are written first.
4. When the relative proportion of an element in a compound is greater than one, the symbol for that element must be followed by a numerical subscript indicating its relative proportion, as in  $MgCl_2$ .
5. When the relative proportion of a polyatomic ion in a compound is greater than one, the symbol for the polyatomic ion must be enclosed by parentheses, followed by the correct numerical subscript, as in  $Al_2(SO_4)_3$ .

In this experiment you will use cut-out models of ions to form neutral compounds. The correct chemical formula and name for each compound will be determined by balancing oxidation numbers and charges.

## OBJECTIVES

In this experiment, you will:

- cut out models of ions
- match the necessary number of ions until the positive and negative oxidation numbers equal 0
- predict the correct formulas for the compounds listed.

## EQUIPMENT

scissors  
pencil and paper  
sheet of ion models  
glue stick

## PROCEDURE

1. Prepare a data table as directed in the Analysis.
2. Cut out each of the "ion" squares from the sheet provided by your teacher.
3. Construct formulas for the following combinations of substances. For example, the formula for a compound containing magnesium and chlorine may be determined in the following way: Place the  $Mg^{2+}$  ion on a piece of paper. Place enough  $Cl^-$  ions alongside the  $Mg^{2+}$  ion to balance the charges. (Positive charges must equal negative charges.)

4. Predict formulas for four additional compounds using the ions listed in this experiment. (See data table for requirements). Use references to determine if your compounds exist.

### Combining Substances:

aluminum and bromine  
sodium and oxygen  
magnesium and sulfur  
aluminum and nitrate ion  
potassium and sulfate ion  
magnesium and chlorine  
ammonium ion and sulfur  
aluminum and oxygen  
magnesium and sulfate ion  
sodium and phosphate ion

# Formulas And Oxidation Numbers Lab Answers

**K Morrison**



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