

- Which event must *always* occur for a chemical reaction to take place?
  - formation of a precipitate
  - formation of a gas
  - effective collisions between reacting particles
  - addition of a catalyst to the reaction system
- In most aqueous reactions as temperature increases, the effectiveness of collisions between reacting particles
  - decreases
  - increases
  - remains the same
- A 5.0-gram sample of zinc and a 50.-milliliter sample of hydrochloric acid are used in a chemical reaction. Which combination of these samples has the fastest reaction rate?
  - a zinc strip and 1.0 M HCl(aq)
  - a zinc strip and 3.0 M HCl(aq)
  - zinc powder and 1.0 M HCl(aq)
  - zinc powder and 3.0 M HCl(aq)
- A 1.0-gram piece of zinc reacts with 5 milliliters of HCl(aq). Which of these conditions of concentration and temperature would produce the greatest rate of reaction?
  - 1.0 M HCl(aq) at 20.°C
  - 1.0 M HCl(aq) at 40.°C
  - 2.0 M HCl(aq) at 20.°C
  - 2.0 M HCl(aq) at 40.°C
- At STP, which 4.0-gram zinc sample will react fastest with dilute hydrochloric acid?
  - lump
  - bar
  - powdered
  - sheet metal
- Given the balanced equation representing a reaction:
 
$$\text{Zn}(s) + 2\text{HCl}(aq) \rightarrow \text{H}_2(g) + \text{ZnCl}_2(aq)$$

Which set of reaction conditions produces  $\text{H}_2(g)$  at the fastest rate?

  - a 1.0-g lump of Zn(s) in 50. mL of 0.5 M HCl(aq) at 20.°C
  - a 1.0-g lump of Zn(s) in 50. mL of 0.5 M HCl(aq) at 30.°C
  - 1.0 g of powdered Zn(s) in 50. mL of 1.0 M HCl(aq) at 20.°C
  - 1.0 g of powdered Zn(s) in 50. mL of 1.0 M HCl(aq) at 30.°C
- For a given reaction, adding a catalyst increases the rate of the reaction by
  - providing an alternate reaction pathway that has a higher activation energy
  - providing an alternate reaction pathway that has a lower activation energy
  - using the same reaction pathway and increasing the activation energy
  - using the same reaction pathway and decreasing the activation energy

- Given the balanced equation representing a reaction:
 
$$\text{CH}_4(g) + 2\text{O}_2(g) \rightarrow 2\text{H}_2\text{O}(g) + \text{CO}_2(g) + \text{heat}$$

Which statement is true about energy in this reaction?

  - The reaction is exothermic because it releases heat.
  - The reaction is exothermic because it absorbs heat.
  - The reaction is endothermic because it releases heat.
  - The reaction is endothermic because it absorbs heat.

- Which balanced equation represents an endothermic reaction?
  - $\text{C}(s) + \text{O}_2(g) \rightarrow \text{CO}_2(g)$
  - $\text{CH}_4(g) + 2\text{O}_2(g) \rightarrow \text{CO}_2(g) + 2\text{H}_2\text{O}(l)$
  - $\text{N}_2(g) + 3\text{H}_2(g) \rightarrow 2\text{NH}_3(g)$
  - $\text{N}_2(g) + \text{O}_2(g) \rightarrow 2\text{NO}(g)$
- Given the balanced equation:
 
$$\text{KNO}_3(s) + 34.89 \text{ kJ} \xrightarrow{\text{H}_2\text{O}} \text{K}^+(aq) + \text{NO}_3^-(aq)$$

Which statement best describes this process?

- It is endothermic and entropy increases.
  - It is endothermic and entropy decreases.
  - It is exothermic and entropy increases.
  - It is exothermic and entropy decreases.
- Which change is exothermic?
    - freezing of water
    - melting of iron
    - vaporization of ethanol
    - sublimation of iodine
  - Which reaction releases the greatest amount of energy per 2 moles of product?
    - $2\text{CO}(g) + \text{O}_2(g) \rightarrow 2\text{CO}_2(g)$
    - $4\text{Al}(s) + 3\text{O}_2(g) \rightarrow 2\text{Al}_2\text{O}_3(s)$
    - $2\text{H}_2(g) + \text{O}_2(g) \rightarrow 2\text{H}_2\text{O}(g)$
    - $\text{N}_2(g) + 3\text{H}_2(g) \rightarrow 2\text{NH}_3(g)$
  - Given the balanced equation representing a reaction at 101.3 kPa and 298 K:
 
$$\text{N}_2(g) + 3\text{H}_2(g) \rightarrow 2\text{NH}_3(g) + 91.8 \text{ kJ}$$

Which statement is true about this reaction?

    - It is exothermic and  $\Delta H$  equals  $-91.8 \text{ kJ}$ .
    - It is exothermic and  $\Delta H$  equals  $+91.8 \text{ kJ}$ .
    - It is endothermic and  $\Delta H$  equals  $-91.8 \text{ kJ}$ .
    - It is endothermic and  $\Delta H$  equals  $+91.8 \text{ kJ}$ .
  - Which expression represents the  $\Delta H$  for a chemical reaction in terms of the potential energy, PE, of its products and reactants?
    - PE of products + PE of reactants
    - PE of products – PE of reactants
    - PE of products = PE of reactants
    - PE of products + PE of reactants

# Kinetics And Equilibrium Regents Questions Answers

**Ensheng Dong**



**Kinetics And Equilibrium Regents Questions Answers:**

## Reviewing **Kinetics And Equilibrium Regents Questions Answers**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Kinetics And Equilibrium Regents Questions Answers**," an enthralling opus penned by a highly acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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