

In Table 4 below,  $[\text{Fe}^{3+}]$  after mixing is obtained by using  $M_1V_1 = M_2V_2$ . This data is to be transferred to Table 5 below as the initial  $[\text{Fe}^{3+}]$  (column 2).

Similarly in Table 4 of Data sheet  $[\text{SCN}^-]$  after mixing obtained using  $M_1V_1 = M_2V_2$ . This data is then to be transferred to Table 5 below as the initial  $[\text{SCN}^-]$  (column 3).

Table 4. Initial Concentration Data

| Test Tube | $[\text{Fe}^{3+}]$<br>Before<br>Mixing | $[\text{Fe}^{3+}]$<br>After Mixing <sup>1</sup><br>(Initial $[\text{Fe}^{3+}]$ ) | $[\text{SCN}^-]$<br>Before<br>Mixing | $[\text{SCN}^-]$<br>After Mixing <sup>1</sup><br>(Initial $[\text{SCN}^-]$ ) |
|-----------|--|--|--------------------------------------|--|
| F         | 0.01                                   | 0.0025M  | 0.00113                              | $2.825 \times 10^{-4}$   |
| G         | 0.01                                   | 0.0025M  | 0.00113                              | $5.65 \times 10^{-4}$  |
| H         | 0.01                                   | 0.005M   | 0.00113                              | $2.825 \times 10^{-4}$   |
| I         | 0.01                                   | 0.005M   | 0.00113                              | $1.13 \times 10^{-4}$  |

(Show calculations for initial concentration data on separate sheet)

In Table 5 below, Equilibrium  $[\text{FeSCN}^{2+}]$  is obtained from the Calibration line using absorbance values of solutions F through I.

Equilibrium  $[\text{Fe}^{3+}] = \text{Initial } [\text{Fe}^{3+}] - \text{Equilibrium } [\text{FeSCN}^{2+}]$

Equilibrium  $[\text{SCN}^-] = \text{Initial } [\text{SCN}^-] - \text{Equilibrium } [\text{FeSCN}^{2+}]$

Table 5. Equilibrium Constant Calculation

| Test Tube | Initial $[\text{Fe}^{3+}]$ | Initial $[\text{SCN}^-]$ | Equilibrium $[\text{FeSCN}^{2+}]$<br>(from graph) | Equilibrium $[\text{Fe}^{3+}]$ | Equilibrium $[\text{SCN}^-]$ | $K_c$ |
|-----------|----------------------------|--------------------------|---|--------------------------------|------------------------------|-------|
| F         | 0.0025                     | $2.825 \times 10^{-4}$   |   |                                |                              |       |
| G         | 0.0025                     | $5.65 \times 10^{-4}$    |   |                                |                              |       |
| H         | 0.005                      | $2.825 \times 10^{-4}$   |   |                                |                              |       |
| I         | 0.005                      | $1.13 \times 10^{-4}$    |   |                                |                              |       |

(Show calculations for Equilibrium Constant on separate sheet)

Average Equilibrium Constant \_\_\_\_\_

# Determining Equilibrium Constant Lab Report Answers

**Yan Bai**



**Determining Equilibrium Constant Lab Report Answers:**

## **Determining Equilibrium Constant Lab Report Answers** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Determining Equilibrium Constant Lab Report Answers**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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