

Stoichiometry Worksheet # 1

1) Glucose is used as a source of energy by the human body. The overall reaction in the body is



Calculate the number of moles of oxygen needed to oxidize 12.5 g of $\text{C}_6\text{H}_{12}\text{O}_6$.

2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.



If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia (NH_3) can be produced?

3) Assume that in the decomposition of potassium chlorate, KClO_3 ,



80.5 moles of O_2 form. How many grams of potassium chlorite, the other product would be formed?

4) In a single displacement reaction,



9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

5) The compound "Cisplatin" $\text{PtCl}_2(\text{NH}_3)_2$, has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction



A. How many moles of "cisplatin" can be produced from 2.50 g K_2PtCl_4 ?

B. How many moles of NH_3 would be needed if 3.8 moles of K_2PtCl_4 were present?

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Georg G. Behrens



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Table of Contents Stoichiometry Chapter 12 Study For Content Mastery Answers

1. Understanding the eBook Stoichiometry Chapter 12 Study For Content Mastery Answers
 - The Rise of Digital Reading Stoichiometry Chapter 12 Study For Content Mastery Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Chapter 12 Study For Content Mastery Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Chapter 12 Study For Content Mastery Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Chapter 12 Study For Content Mastery Answers
 - Personalized Recommendations
 - Stoichiometry Chapter 12 Study For Content Mastery Answers User Reviews and Ratings

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- Stoichiometry Chapter 12 Study For Content Mastery Answers and Bestseller Lists
- 5. Accessing Stoichiometry Chapter 12 Study For Content Mastery Answers Free and Paid eBooks
 - Stoichiometry Chapter 12 Study For Content Mastery Answers Public Domain eBooks
 - Stoichiometry Chapter 12 Study For Content Mastery Answers eBook Subscription Services
 - Stoichiometry Chapter 12 Study For Content Mastery Answers Budget-Friendly Options
- 6. Navigating Stoichiometry Chapter 12 Study For Content Mastery Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Stoichiometry Chapter 12 Study For Content Mastery Answers Compatibility with Devices
 - Stoichiometry Chapter 12 Study For Content Mastery Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Chapter 12 Study For Content Mastery Answers
 - Highlighting and Note-Taking Stoichiometry Chapter 12 Study For Content Mastery Answers
 - Interactive Elements Stoichiometry Chapter 12 Study For Content Mastery Answers
- 8. Staying Engaged with Stoichiometry Chapter 12 Study For Content Mastery Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Chapter 12 Study For Content Mastery Answers
- 9. Balancing eBooks and Physical Books Stoichiometry Chapter 12 Study For Content Mastery Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Chapter 12 Study For Content Mastery Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Stoichiometry Chapter 12 Study For Content Mastery Answers
 - Setting Reading Goals Stoichiometry Chapter 12 Study For Content Mastery Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Stoichiometry Chapter 12 Study For Content Mastery Answers
 - Fact-Checking eBook Content of Stoichiometry Chapter 12 Study For Content Mastery Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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