

Chapter 10 test**Matching**

Match each item with the correct statement below.

- | | |
|----------------------------|--------------------------------------|
| a. representative particle | d. percent composition |
| b. mole | e. standard temperature and pressure |
| c. Avogadro's number | f. empirical formula |

- _____ 1. the number of representative particles of a substance present in 1 mole of that substance
- _____ 2. an atom, an ion, or a molecule, depending upon the way a substance commonly exists
- _____ 3. the SI unit used to measure amount of substance
- _____ 4. 0°C and 1 atm
- _____ 5. the percent by mass of each element in a compound
- _____ 6. the smallest whole number ratio of the atoms in a compound

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- _____ 7. What SI unit is used to measure the number of representative particles in a substance?
- | | |
|-------------|-----------|
| a. kilogram | c. kelvin |
| b. ampere | d. mole |
- _____ 8. How many hydrogen atoms are in 5 molecules of isopropyl alcohol, C₃H₇O?
- | | |
|-------------------------------------|--------------------------------------|
| a. $5 \times (6.02 \times 10^{23})$ | c. 35 |
| b. 5 | d. $35 \times (6.02 \times 10^{23})$ |
- _____ 9. All of the following are equal to Avogadro's number EXCEPT _____
- | |
|--|
| a. the number of atoms of bromine in 1 mol Br ₂ |
| b. the number of atoms of gold in 1 mol Au |
| c. the number of molecules of nitrogen in 1 mol N ₂ |
| d. the number of molecules of carbon monoxide in 1 mol CO |
- _____ 10. How many moles of tungsten atoms are in 4.8×10^{23} atoms of tungsten?
- | | |
|-------------------------------|-------------------------------|
| a. 8.0×10^{-2} moles | c. 1.3×10^{-4} moles |
| b. 8.0×10^1 moles | d. 1.3×10^{-2} moles |
- _____ 11. How many atoms are in 0.075 mol of titanium?
- | | |
|-------------------------|-------------------------|
| a. 1.2×10^{23} | c. 6.4×10^2 |
| b. 2.2×10^{24} | d. 4.5×10^{22} |
- _____ 12. How many molecules are in 2.10 mol CO₂?
- | | |
|------------------------------------|-------------------------------------|
| a. 2.53×10^{24} molecules | c. 3.49×10^{-24} molecules |
| b. 3.79×10^{24} molecules | d. 1.26×10^{24} molecules |

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JR Anderson



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