

Chapter 6 Chemical Bonds

Section 6.4 The Structure of Metals**(pages 176–181)**

This section discusses metallic bonds and the properties of metals. It also explains how the properties of an alloy are controlled.

Reading Strategy (page 176)

Relating Cause and Effect As you read, complete the concept map to relate the structure of metals to their properties. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.



- Circle the letter of the metal with the highest melting point.
 - gold
 - vanadium
 - titanium
 - tungsten
- Is the following sentence true or false? The properties of a metal are related to bonds within the metal. true

Metallic Bonds (pages 176–177)

- Describe a metallic bond. A metallic bond is the attraction between a metal cation and the cation and the shared electrons that surround it.
- The cations in a metal form a lattice. What holds the lattice in place? The lattice is held in place by strong metallic bonds between the cations and the surrounding valence electrons.
- Is the following sentence true or false? The more valence electrons a metal has, the stronger its metallic bonds will be. true

Explaining Properties of Metals (page 177)

- Some of the properties of metals can be explained by the mobility of the electrons within a metal lattice.
- Name two important properties of metals that can be explained by metallic bonding.
 - Malleability
 - Ability to conduct an electric current

Alloys (pages 178–181)

- Circle the letter of the percentage of gold in jewelry that is labeled 18-karat gold.
 - 18 percent
 - 50 percent
 - 75 percent
 - 100 percent

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Table of Contents Quick Switch Answer Key

1. Understanding the eBook Quick Switch Answer Key
 - The Rise of Digital Reading Quick Switch Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Quick Switch Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Quick Switch Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Quick Switch Answer Key
 - Personalized Recommendations
 - Quick Switch Answer Key User Reviews and Ratings
 - Quick Switch Answer Key and Bestseller Lists
5. Accessing Quick Switch Answer Key Free and Paid eBooks
 - Quick Switch Answer Key Public Domain eBooks
 - Quick Switch Answer Key eBook Subscription Services
 - Quick Switch Answer Key Budget-Friendly Options
6. Navigating Quick Switch Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Quick Switch Answer Key Compatibility with Devices
 - Quick Switch Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Quick Switch Answer Key
 - Highlighting and Note-Taking Quick Switch Answer Key
 - Interactive Elements Quick Switch Answer Key
8. Staying Engaged with Quick Switch Answer Key

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Quick Switch Answer Key
9. Balancing eBooks and Physical Books Quick Switch Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Quick Switch Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Quick Switch Answer Key
- Setting Reading Goals Quick Switch Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Quick Switch Answer Key
- Fact-Checking eBook Content of Quick Switch Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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