

Empirical and Molecular Formula Worksheet ANSWER KEY

Write the empirical formula for the following compounds.

- 1) C_6H_6 **CH**
- 6) C_8H_{12} **C_4H_6**
- 7) WO_2 **WO_2**
- 8) $C_2H_6O_2$ **CH_3O**
- 9) $X_{10}Y_{12}$ **X_5Y_6**
- 6) A compound with an empirical formula of C_2OH_4 and a molar mass of 88 grams per mole. What is the molecular formula of this compound? **$C_8O_4H_{16}$** 
- 7) A compound with an empirical formula of C_4H_6O and a molar mass of 136 grams per mole. What is the molecular formula of this compound? **$C_8H_{12}O_2$** 
- 8) A compound with an empirical formula of $CFBrO$ and a molar mass of 253.8 grams per mole. What is the molecular formula of this compound? **$C_2F_2Br_2O_2$** 
- 9) A compound with an empirical formula of C_2H_3N and a molar mass of 46 grams per mole. What is the molecular formula of this compound? **$C_4H_6N_2$** 
- 10) A well-known reagent in analytical chemistry, dimethylglyoxime, has the empirical formula C_2H_5NO . If its molar mass is 116.1 g/mol, what is the molecular formula of the compound? **$C_8H_{16}N_2O_2$** 
12. A certain blue solid contains 36.84% N and 63.16% O. What is the empirical formula of this compound? **The ratios are $N_{1.50}O_{1.50}$. Since 1.50 is not close to a whole number, we multiply *both* subscripts by 2. The empirical formula is thus N_2O_2 . (The name is dinitrogen trioxide.)** 
13. A sample of indium chloride weighing 0.5000 g is found to contain 0.2404 g of chlorine. What is the empirical formula of the indium compound? **$InCl_2$** 
14. An unknown compound was found to have a percent composition as follows: 47.0 % potassium, 14.5 % carbon, and 38.5 % oxygen. What is its empirical formula? If the true molar mass of the compound is 166.22 g/mol, what is its molecular formula? **$K_2C_3O_4$** 
15. Rubbing alcohol was found to contain 60.0 % carbon, 13.4 % hydrogen, and the remaining mass was due to oxygen. What is the empirical formula of rubbing alcohol? **C_3H_8O** 

Molecular Formula Worksheet Answer Key

NARAYAN CHANGDER



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Table of Contents Molecular Formula Worksheet Answer Key

1. Understanding the eBook Molecular Formula Worksheet Answer Key
 - The Rise of Digital Reading Molecular Formula Worksheet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Molecular Formula Worksheet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molecular Formula Worksheet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molecular Formula Worksheet Answer Key

- Personalized Recommendations
 - Molecular Formula Worksheet Answer Key User Reviews and Ratings
 - Molecular Formula Worksheet Answer Key and Bestseller Lists
5. Accessing Molecular Formula Worksheet Answer Key Free and Paid eBooks
- Molecular Formula Worksheet Answer Key Public Domain eBooks
 - Molecular Formula Worksheet Answer Key eBook Subscription Services
 - Molecular Formula Worksheet Answer Key Budget-Friendly Options
6. Navigating Molecular Formula Worksheet Answer Key eBook Formats
- ePub, PDF, MOBI, and More
 - Molecular Formula Worksheet Answer Key Compatibility with Devices
 - Molecular Formula Worksheet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Molecular Formula Worksheet Answer Key
 - Highlighting and Note-Taking Molecular Formula Worksheet Answer Key
 - Interactive Elements Molecular Formula Worksheet Answer Key
8. Staying Engaged with Molecular Formula Worksheet Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molecular Formula Worksheet Answer Key
9. Balancing eBooks and Physical Books Molecular Formula Worksheet Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molecular Formula Worksheet Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Molecular Formula Worksheet Answer Key
- Setting Reading Goals Molecular Formula Worksheet Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Molecular Formula Worksheet Answer Key

- Fact-Checking eBook Content of Molecular Formula Worksheet Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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