

## Question 1: Kinetics of a Reaction and Equilibrium

### 1.1. Determining Reaction Order and Rate Law

- a. Second
- b. First
- c. Third
- d. Fourth
- e. Zero

1.2. Interpret the rate law and determine the overall order of the reaction.

1.3. Calculate the half-life of a first-order reaction. The half-life of a substance is the time required for its concentration to decrease by 50%.

- a. 1
- b. 2
- c. 3
- d. 4
- e. 5
- f. 6
- g. 7
- h. 8

### 1.2. Determining Reaction Order and Equilibrium

- a.  $2x + 3y \rightleftharpoons 4z$
- b.  $2x + 3y \rightleftharpoons 4z + 2w$
- c.  $2x + 3y \rightleftharpoons 4z + 2w + 2v$

1.3. Calculate the equilibrium constant for the reaction of the gases  $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$  at 25°C. The initial concentrations are  $[\text{NO}_2] = 0.100 \text{ M}$  and  $[\text{N}_2\text{O}_4] = 0.000 \text{ M}$ . At equilibrium,  $[\text{NO}_2] = 0.076 \text{ M}$ .

- a.  $1.5 \times 10^{-2}$
- b.  $1.5 \times 10^{-3}$
- c.  $1.5 \times 10^{-4}$
- d.  $1.5 \times 10^{-5}$
- e.  $1.5 \times 10^{-6}$
- f.  $1.5 \times 10^{-7}$
- g.  $1.5 \times 10^{-8}$
- h.  $1.5 \times 10^{-9}$

1.4. Calculate the equilibrium constant for the reaction of the gases  $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$  at 25°C.

### 1.3. Calculating Reaction Free Energy and Entropy

- a.  $1.5 \times 10^3 \text{ J/mol}$

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**Puja Mehta**



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