

Determination of K_a for a Weak Acid

Pre Lab Calculations

*Watch the provided video with example calculations before doing these problems.

Instructions: Watch the video, work these problems and upload your work and answers before your scheduled Zoom lab time. Do your best and use the video. ****MUST show work to receive credit!**

1. The pH of a 0.20 M solution of an unknown monoprotic acid, HA, is 2.12. Determine the value of the K_a for this acid. Start each problem by writing a balanced acid ionization (since calculating K_a) in water reaction and setting up an appropriate ICE table.

Step 1: Acid ionization reaction: $HA(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + A^-(aq)$.

Step 2: Set up an ICE table

	HA	H ₂ O	A ⁻	H ₃ O ⁺
I	0.20M		0	0
C	-x		x	x
E	0.19M		$7.6 \cdot 10^{-3}M$	$7.6 \cdot 10^{-3}M$

Step 3: Determine the $[H_3O^+]_{\text{equil}}$ using the pH.

$$[H_3O^+]_{\text{equil}} = 10^{\text{pH}} = 10^{-2.12} = 7.6 \cdot 10^{-3} M$$

Step 4: Fill in your known values in to the ICE table. Do $C_1V_1 = C_2V_2$ dilutions if necessary, to find concentrations. (Hint: this isn't necessary in question 1, but is necessary if you're mixing your acid with anything else).

Step 5: Write the equilibrium expression using your balanced equation from Step 1. Use equilibrium values in your ICE table to solve for K_a .

$$K_a = \frac{[H_3O^+][A^-]}{[HA]} = \frac{(7.6 \cdot 10^{-3} M)^2}{0.19M} = 3.04 \cdot 10^{-4}$$

2. A solution is prepared by adding 25.0 mL of 0.20 M HA (a different monoprotic acid than in question 1) to 50.0 mL of 0.20 M NaA. The pH of this solution is 3.45. Determine the value of K_a for this acid.

Hint: Follow the steps as outlined in problem 1. Use $C_1V_1 = C_2V_2$ to find the new initial concentration of acid after dilution with the NaA solution. The equation you used in question 1

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Arne Nels Langsjoen



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