

Atomic Structure Review

1. The 3 particles of the atom are:

	Location in Atom?	Size of LOCATION?	Charge?	Mass?
Protons	nucleus	.0001 %	+	1 amu
Neutrons	nucleus	.0001 %	neutral	1 amu
Electrons	cloud	99.999 %	-	.00054 amu

2. The number of protons in one atom of an element determines the atom's ^{properties} ~~atomic #~~, and the number of electrons determines the charge of an element.
3. The atomic number tells you the number of protons in one atom of an element. It also tells you the number of electrons in a neutral atom of that element.
4. The atomic number gives the "identity" of an element as well as its location on the Periodic Table. No two different elements will have the same atomic number.
5. The avg atomic mass of an element is the average mass of an element's naturally occurring atom, or isotopes, taking into account the # of neutrons of each isotope.
6. The mass # of an element is the total number of protons and neutrons in the nucleus of the atom.
7. The mass number is used to calculate the number of neutrons in one atom of an element. In order to calculate the number of neutrons you must subtract the protons from the neutrons.

Atomic Structure Periodic Table Answer Key

M Mosston



Atomic Structure Periodic Table Answer Key:

Atomic Structure Periodic Table Answer Key Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has been apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Atomic Structure Periodic Table Answer Key**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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