

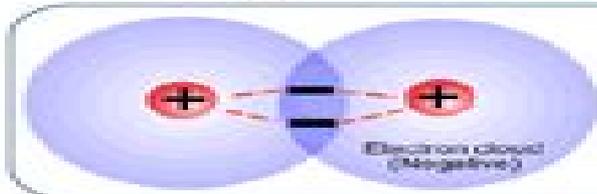
You need to know the content of this sheet. 100%

100% Sheet COVALENT BONDING



What's the **BIG** idea?

Matter
All material in the Universe is made of very small particles.



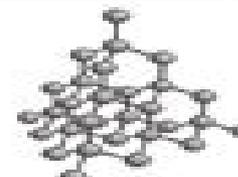
A Covalent bond forms between **2 non-metals** where each atom **shares** its electrons in order to obtain a full outer shell

Simple (small) covalent molecules are mostly gases at room temperature (except water) because they have **weak (easy to break) bonds between the molecules**. The **STRONG COVALENT bonds DO NOT BREAK**



Graphite

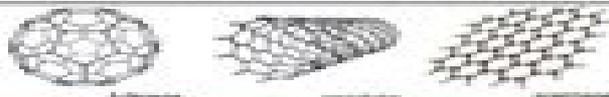
- 3 strong covalent bonds
- 1 weak intermolecular bond between the layers
- layers can slide over each other
- Free (Delocalised) electrons between the layers so it can conduct electricity



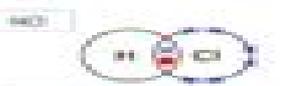
Diamond

- 4 strong covalent bonds so high melting point
- No weak bonds so it is very strong and hard
- No free electrons or charges so it cannot conduct electricity

Other structures of carbon can be man-made. Nanotubes are made by folding a sheet of graphene (which is a single sheet of graphite) these also have delocalised electrons



You **MUST** be able to draw dot and cross diagrams for the molecules of hydrogen, chlorine, oxygen, nitrogen, hydrogen chloride, water, ammonia and methane



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JG Myers



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Table of Contents Covalent Bonding Core Teaching Resources Test Answer

1. Understanding the eBook Covalent Bonding Core Teaching Resources Test Answer
 - The Rise of Digital Reading Covalent Bonding Core Teaching Resources Test Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Bonding Core Teaching Resources Test Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Covalent Bonding Core Teaching Resources Test Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Bonding Core Teaching Resources Test Answer
 - Personalized Recommendations

- Covalent Bonding Core Teaching Resources Test Answer User Reviews and Ratings
- Covalent Bonding Core Teaching Resources Test Answer and Bestseller Lists
- 5. Accessing Covalent Bonding Core Teaching Resources Test Answer Free and Paid eBooks
 - Covalent Bonding Core Teaching Resources Test Answer Public Domain eBooks
 - Covalent Bonding Core Teaching Resources Test Answer eBook Subscription Services
 - Covalent Bonding Core Teaching Resources Test Answer Budget-Friendly Options
- 6. Navigating Covalent Bonding Core Teaching Resources Test Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Covalent Bonding Core Teaching Resources Test Answer Compatibility with Devices
 - Covalent Bonding Core Teaching Resources Test Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Covalent Bonding Core Teaching Resources Test Answer
 - Highlighting and Note-Taking Covalent Bonding Core Teaching Resources Test Answer
 - Interactive Elements Covalent Bonding Core Teaching Resources Test Answer
- 8. Staying Engaged with Covalent Bonding Core Teaching Resources Test Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Covalent Bonding Core Teaching Resources Test Answer
- 9. Balancing eBooks and Physical Books Covalent Bonding Core Teaching Resources Test Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Covalent Bonding Core Teaching Resources Test Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Covalent Bonding Core Teaching Resources Test Answer
 - Setting Reading Goals Covalent Bonding Core Teaching Resources Test Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Covalent Bonding Core Teaching Resources Test Answer
 - Fact-Checking eBook Content of Covalent Bonding Core Teaching Resources Test Answer

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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