

Level 1

- The vapour pressure of a given liquid will decrease if :
 - surface area of liquid is decreased
 - the volume of liquid in the container is decreased
 - the volume of the vapour phase is increased
 - the temperature is decreased
- The normal boiling point of water is 373 K. Vapour pressure of water at temperature T is 19 mm Hg. If enthalpy of vaporisation is 40.67 kJ/mol, then temperature T would be
(Use : $\log 2 = 0.3$, $R : 8.3 \text{ JK}^{-1} \text{ mol}^{-1}$):
 - 250 K
 - 291.4 K
 - 230 K
 - 290 K
- A sample of liquid H_2O of mass 18.0 g is injected into an evacuated 7.6 L flask maintained at 27.0°C . If vapour pressure of H_2O at 27°C is 24.63 mm Hg. What weight percentage of the water will be vapourised when the system comes to equilibrium? Assume water vapours behaves as an ideal gas. The volume occupied by the liquid water is negligible compared to the volume of the container:
 - 1%
 - 10%
 - 18%
 - 20%
- Raoult's law is obeyed by each constituent of a binary liquid solution when:
 - the forces of attractions between like molecules are greater than those between unlike molecules
 - the forces of attractions between like molecules are smaller than those between unlike molecules
 - the forces of attractions between like molecules are identical with those between unlike molecules
 - the volume occupied by unlike molecules are different
- For a binary ideal liquid solution, the total pressure of the solution is given as:
 - $P_{\text{total}} = P_A^* + (P_A^* - P_B^*) X_B$
 - $P_{\text{total}} = P_B^* + (P_A^* - P_B^*) X_A$
 - $P_{\text{total}} = P_B^* + (P_B^* - P_A^*) X_A$
 - $P_{\text{total}} = P_B^* + (P_B^* - P_A^*) X_B$
- For an ideal binary liquid solution with $P_A^* > P_B^*$, which relation between X_A (mole fraction of A in liquid phase) and Y_A (mole fraction of A in vapour phase) is correct ?
 - $Y_A < Y_B$
 - $X_A > X_B$
 - $\frac{Y_A}{Y_B} > \frac{X_A}{X_B}$
 - $\frac{Y_A}{Y_B} < \frac{X_A}{X_B}$
- An ideal solution is formed by mixing two volatile liquids A and B. X_A and X_B are the mole fractions of A and B respectively in the solution and Y_A and Y_B are the mole fractions of A and B respectively in the vapour phase. A plot of $\frac{1}{Y_A}$ along y-axis against $\frac{1}{X_A}$ along x-axis gives a straight line. What is the slope of the straight line ?
 - $\frac{P_B^*}{P_A^*}$
 - $\frac{P_A^*}{P_B^*}$
 - $P_B^* - P_A^*$
 - $P_A^* - P_B^*$

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Xiaolong Qi



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N Awasthi Solutions Pdf Book Review: Unveiling the Magic of Language

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