



NATIONAL
EXAM CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2
(CHEMISTRY)
EXEMPLAR

NAME: _____
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QUESTION 3

PHYSICAL SCIENCES PAPER 2

INTERMOLECULAR FORCES

GRADE 11

NOVEMBER 2020

QUESTION NO. _____ INITIAL ANSWERS _____

QUESTION 3 starts on P209 page 1.

3.1 The boiling points in the table below were obtained during an investigation into the boiling points of pure 1-hexanol. The table below lists the boiling points of pure 1-hexanol and pure 1-hexanol with 10% water.

Compound	Boiling point (°C)	Boiling point (°C)
A	158	158
B	158	158

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 In which column are compounds A and B in (1) and (2) lists of boiling points? (2)
- 3.1.3 Name the type of intermolecular force that water shows between molecules of pure compounds A and B due to the polar nature of their molecules. (2)
- 3.1.4 What type of intermolecular force is responsible for the boiling point of pure 1-hexanol? (2)
- 3.1.5 What compound (A or B) will have the higher vapour pressure? (2)
- 3.1.6 Give a reason for your answer to 3.1.5. (2)

Compound B (10% of water in the table) has a smaller boiling point than pure compounds A and B but has a smaller vapour boiling point of 154 °C.

- 3.1.7 Explain why the boiling point of compound B is smaller than that of compound A and B due to the nature of the 1-hexanol and water intermolecular forces. (2)

3.2 Additive D is added to each mixture of pure compounds A and B to lower its boiling point down to that of the other mixture.



- 3.2.1 In which beaker (1) or (2), will the water be more? (2)
- 3.2.2 Explain the reason for (QUESTION 3.2.1) based on the nature of the intermolecular forces. (2)

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Physical Science Paper 2 Intermolecular Forces Investigation

S Ashworth



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