

LAB 06 – DOUBLE DISPLACEMENT REACTIONS

Discussion - Double Replacement Reactions (Refer to class notes from Chapter 7)

All reactions in this part of the experiment are double replacement reactions.

Double replacement reactions, sometimes called "double displacement", "exchange reactions" or "trading partners" are very common. Generally, they occur when ionic compounds are dissolved in aqueous (water) solution. They have the general form:



where A and B are positive ions and X and Y are negative ions.

A reaction will take place if one of the following is formed in the reaction:

- a **precipitate** (what you see is a solid material that clouds the solution).
- a **gas** (you see lots of bubbles coming out of solution).
- water** (You will feel the test tube get warm as heat is given off by the reaction).

These are known as the **Driving Force** of the reaction.

If the reaction produces neither a precipitate, nor a gas, nor a molecular substance, then **no reaction** has taken place. If a reaction does take place, then moving from a balanced molecular equation to a net ionic equation allows us to determine exactly what the chemistry was in the reaction, eliminating the non-reacting (spectator) ions and/or molecules.

The following double displacement reactions were carried out in lab and the listed observations were made.

Reagents	Observations
1. $\text{NaCl}_{(aq)} + \text{AgNO}_{3(aq)}$	<i>White precipitate formed</i>
2. $\text{Na}_2\text{CO}_{3(aq)} + \text{HCl}_{(aq)}$	<i>Bubbles form, solution got warm</i>
3. $\text{NaOH}_{(aq)} + \text{HCl}_{(aq)}$	<i>Solution got warm</i>
4. $\text{Na}_2\text{CO}_{3(aq)} + \text{CaCl}_{2(aq)}$	<i>White precipitate formed</i>
5. $\text{FeCl}_{3(aq)} + \text{NH}_4\text{OH}_{(aq)}$	<i>Reddish brown precipitate formed</i>
6. $\text{CuSO}_{4(aq)} + \text{Zn}(\text{NO}_3)_2(aq)$	<i>No Reaction</i>

Double Displacement Reaction Lab Answers

Kathy Dodds Tyner



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Lab Experiments in Introductory Chemistry Phil Reedy, Donald J. Wink, Sharon Fetzer-Gislason, 2003-03-21 The manual contains laboratory experiments written specifically for the prep chem lab as well as for the general chemistry course Available as a complete manual or custom published at <http://custompub.whfreeman.com> **Lab Manual Science Class 10** Neena Sinha, R. Rangarajan, Rajesh Kumar, These Lab Manuals provide complete information on all the experiments listed in the latest CBSE syllabus The various objectives materials required procedures inferences etc have been given in a step by step manner Carefully framed MCQs and short answers type questions given at the end of the experiments help the students prepare for viva voce *LK-Science-HB-10-R* R Rangarajan, RP Manchanda, Rajesh Kumar, LK Science HB 10 R **E3 Chemistry Review Book - 2018 Home Edition (Answer Key Included)** Effiong Eyo, 2017-10-20 With Answer Key to All Questions Chemistry students and homeschoolers Go beyond just passing Enhance your understanding of chemistry and get higher marks on homework quizzes tests and the regents exam with E3 Chemistry Review Book 2018 With E3 Chemistry Review Book students will get clean clear engaging exciting and easy to understand high school chemistry concepts with emphasis on New York State Regents Chemistry the Physical Setting Easy to read format to help students easily remember key and must know chemistry materials Several example problems with solutions to study and follow Several practice multiple choice and short answer questions at the end of each lesson to test understanding of the materials 12 topics of Regents question sets and 3 most recent Regents exams to practice and prep for any Regents Exam This is the Home Edition of the book Also available in School Edition ISBN 978 197836229 The Home Edition contains an answer key section Teachers who want to recommend our Review Book to their students should recommend the Home Edition Students and parents whose school is not using the Review Book as instructional material as well as homeschoolers should buy the Home Edition The School Edition does not have answer key in the book A separate answer key booklet is provided to teachers with a class order of the book Whether you are using the school or Home Edition our E3 Chemistry Review Book makes a great supplemental instructional and test prep resource that can be used from the beginning to the end of the school year PLEASE NOTE Although reading contents in both the school and home editions are identical there are slight differences in question numbers choices and pages between the two editions Students whose school is using the Review Book as instructional material SHOULD NOT buy the Home Edition Also available in paperback print *Illustrated Guide to Home Chemistry Experiments* Robert Bruce Thompson, 2012-02-17 For students DIY hobbyists and science buffs who can no longer get real chemistry sets this one of a kind guide explains how to set up and use a home chemistry lab with step by step instructions for conducting experiments in basic chemistry not just to make pretty colors and stinky smells but to learn how to do real lab work Purify alcohol by distillation Produce hydrogen and oxygen gas by electrolysis Smelt metallic copper from copper ore you make yourself Analyze the makeup of seawater bone and other common substances Synthesize oil of wintergreen from

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Chemistry Eugene LeMay, Jr.,Herbert Beall,Karen M. Robblee,Douglas C. Brower,Prentice-Hall Staff,2002-02

Whole-class Inquiry Dennis W. Smithenry,Joan Gallager-Bolos,2009 In response to requests from science education professionals this is the perfect vehicle for implementing and assessing this concept of whole class inquiry in your classroom This is a must have package for preservice and inservice middle and high school science teachers

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Chemistry in the Laboratory James M. Postma,Julian L. Roberts,J. Leland Hollenberg,2009-08-07 For nearly 40 years Chemistry in the Laboratory has been meeting the needs of teachers and students This new edition builds on that legacy while addressing cutting edge trends in the chemistry laboratory including forensic chemistry and environmental and green chemistry As always the new edition of Chemistry in the Laboratory offers precise easy to follow instructions helpful illustrations and an emphasis throughout on laboratory safety Again throughout a Consider This feature encourages users to expand the principles of the experiment into interesting applications open ended experiments or unexplored corners Most experiments in the manual can be completed in one lab session but some can be linked or extended for a multi lab project

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