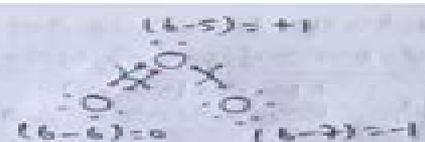


$0 \rightarrow 1 \rightarrow 0_2$



→ Stability of same molecules of diff sp

- ① More stable due to symmetrical charge distribution.
- ② More stable due to absence of charge.
- ③ More stable due to -ve charge present on high EN element.

# Avg. formal charge & Avg. X-O bond order in oxo-anions:-

$$\text{Avg. f.c.} = \frac{\text{total charge present on O-atom}}{\text{No. of O-atom where -ve charge can be placed}}$$

$$\text{Avg. X-O bond order} = \frac{\text{total no. of X-O bonds}}{\text{no. of X-O pairs participate in resonance}}$$



F.C on 'O' → -2/3  
X-O bond order → 4/3

-2/4  
6/4

TRICK	FC
= O →	0
- O →	-1
≡ O or	+1
= O -	

# Covalent bond

↳ Formed by the equal sharing of  $e^-$

\* **Covalency**:- Total no. of covalent bond ( $\sigma$  &  $\pi$ ) formed by an atom in a molecule.

VALENCY (combining capacity)

① **Covalency** → (Total no. of  $\sigma$  &  $\pi$  bonds)

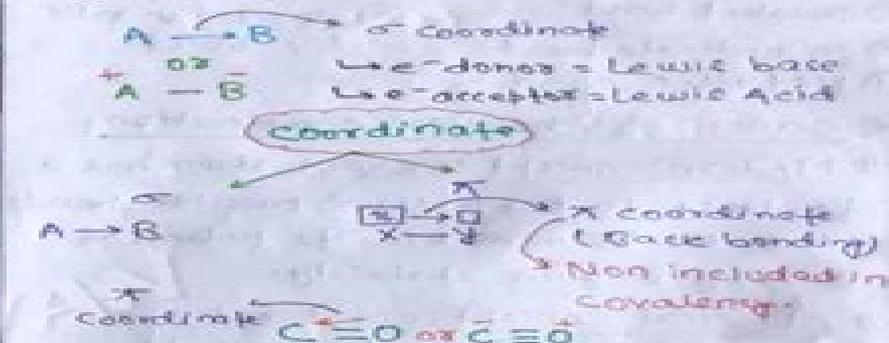
② **Electrovalency** → Combining capacity in ionic comp. E.g.  $\text{Na}^+\text{Cl}^- \rightarrow \text{Na}^+ = +1$   
 $\text{Cl}^- = -1$

NOTE:-

- ① Max covalency of 2nd period element can not be more than 4 due to absence of d-orbitals.
- ② 3rd period elements can extend their covalency beyond 4 by using vacant 3d orbitals.
- ③ 'F' can show only 1 covalency.

# Coordinate bond

↳ Formed by the unequal sharing of  $e^-$   
↳ formed by b/w  $e^-$  donor &  $e^-$  acceptor



# Species in which  $\sigma$  coordinate bond (dative bond) is present:

- ①  $\text{O}_3 \rightarrow \text{O}=\text{O} \rightarrow \text{O}$  or  $\text{O}=\text{O}^+-\text{O}^-$
- ②  $\text{NH}_4^+ \rightarrow [\text{NH}_3 \rightarrow \text{H}^+]$
- ③  $\text{H}_3\text{O}^+ \rightarrow [\text{H}_2\text{O} \rightarrow \text{H}^+]$
- ④  $\text{NO}_3^- \rightarrow \text{O}=\text{N}^+=\text{O} \text{ or } \text{O}^--\text{N}^+=\text{O}$       (2/3)  
(1/3)
- ⑤  $\text{PCl}_6^- \rightarrow (\text{Cl}^- \rightarrow \text{PCl}_5)$
- ⑥  $\text{BF}_4^- \rightarrow \text{F}^- \rightarrow \text{BF}_3$
- ⑦  $\text{NH}_3 \cdot \text{BF}_3 \rightarrow \text{H}_3\text{N} \rightarrow \text{BF}_3$
- ⑧  $\text{PH}_4^+ [\text{H}_2\text{P} \rightarrow \text{H}^+]$
- ⑨  $\text{B(OH)}_4^- \rightarrow \text{B(OH)}_3 \cdot \text{OH}^-$

# Engineering Chemistry Full Notes

**Sabine Zange**



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