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**Principles of Physical Chemistry** B.R. Puri, L.R. Sharma, M.S. Pathania, 2008-03-01      **Chemistry for Engineers** B. K. Ambasta, 2006      *A Textbook of Physical Chemistry, 6th Edition* Sharma K.K. & Sharma L.K., 2016 A Textbook of Physical Chemistry      [A Textbook of Physical Chemistry - Volume 1](#) Mandeep Dalal, 2018-01-01 An advanced level textbook of physical chemistry for the graduate B Sc and postgraduate M Sc students of Indian and foreign universities This book is a part of four volume series entitled A Textbook of Physical Chemistry Volume I II III IV

**CONTENTS**

**Chapter 1 Quantum Mechanics I** Postulates of quantum mechanics Derivation of Schrodinger wave equation Max Born interpretation of wave functions The Heisenberg s uncertainty principle Quantum mechanical operators and their commutation relations Hermitian operators elementary ideas quantum mechanical operator for linear momentum angular momentum and energy as Hermitian operator The average value of the square of Hermitian operators Commuting operators and uncertainty principle  $x$   $E$  Schrodinger wave equation for a particle in one dimensional box Evaluation of average position average momentum and determination of uncertainty in position and momentum and hence Heisenberg s uncertainty principle Pictorial representation of the wave equation of a particle in one dimensional box and its influence on the kinetic energy of the particle in each successive quantum level Lowest energy of the particle

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Schrodinger wave equation for three dimensional Rigid rotator Energy of rigid rotator Space quantization Schrodinger wave equation for hydrogen atom separation of variable in polar spherical coordinates and its solution Principle azimuthal and magnetic quantum numbers and the magnitude of their values Probability distribution function Radial distribution function Shape of atomic orbitals s p Law of mass action and its thermodynamic derivation Third law of thermodynamics Nernst heat theorem determination of absolute entropy unattainability of absolute zero and its limitation Phase diagram for two completely miscible components systems Eutectic systems Calculation of eutectic point Systems forming solid compounds Ax By with congruent and incongruent melting points Phase diagram and thermodynamic treatment of solid solutions Chapter 7 Chemical Dynamics II Chain reactions hydrogen bromine reaction pyrolysis of acetaldehyde decomposition of ethane Photochemical reactions hydrogen bromine General treatment of chain reactions ortho para hydrogen conversion and hydrogen bromine reactions Apparent activation energy of chain reactions Chain length Rice Herzfeld mechanism of organic molecules decomposition acetaldehyde Branching chain reactions and explosions H<sub>2</sub> O<sub>2</sub> reaction Kinetics of one intermediate enzymatic reaction Michaelis Menton treatment Evaluation of Michaelis s constant for enzyme substrate binding by Lineweaver Burk plot and Eadie Hofstae methods Competitive and non competitive inhibition Chapter 8 Electrochemistry II Ion Transport in Solutions Ionic movement under the influence of an electric field Mobility of ions Ionic drift velocity and its relation with current density Einstein relation between the absolute mobility and diffusion coefficient The Stokes Einstein relation The Nernst Einstein equation Walden s rule The Rate process approach to ionic migration The Rate process equation for equivalent conductivity Total driving force for ionic transport Nernst Planck Flux equation Ionic drift and diffusion potential the Onsager phenomenological equations The basic equation for the diffusion Planck Henderson equation for the diffusion potential

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