

Preparation of Soluble Salts

**By Neutralization**

25.0cm<sup>3</sup> acid, as standard solution, is placed in conical flask using pipette. Add few drops of indicator & titrate with alkali from burette until indicator changes colour, showing all acid has just reacted. Volume of alkali added is measured. Prepare new 25.0cm<sup>3</sup> acid again with pipette & add same volume of alkali as before to prevent excess alkali/acid because both reactant & product are aqueous. Next, the product is evaporated to dryness to obtain the salt.



1. Measure accurately (25.0 cm<sup>3</sup>) of acid into a conical flask. Add a few drops of indicator. It is best when the indicator is added to the acid.
2. The burette is now prepared (25.0 cm<sup>3</sup> of alkali) and is placed in a stand. Read the volume. The volume of alkali added is given by the difference between the volume of alkali added and the volume of alkali remaining.
3. The product solution is prepared by evaporating most of the water and crystallizing the salt.

**By Reacting Metal with Acid**



1. Add the pieces of metal to the acid slowly to prevent splashing. Add a volume of acid to cover the metal. The reaction will stop if there is excess acid present.



2. Remove the metal strip slowly. The reaction will stop if there is excess acid present.



3. Heat the contents of your beaker over an evaporating dish until the liquid is nearly dry. The products of the reaction are a solid residue. Heat the residue to dry.



4. Heat the acid solution slowly and try to heat to evaporating most of the water. The products of the reaction are a solid residue. Heat the residue to dry.

Only metals like zinc and magnesium, which moderately react with dilute acids, are used.

E.g. Reacting Zn with H<sub>2</sub>SO<sub>4</sub> to prepare ZnSO<sub>4</sub>



Zn is added to H<sub>2</sub>SO<sub>4</sub> until in excess to ensure no more H<sub>2</sub>SO<sub>4</sub> is present. Then the mixture is filtered off to separate Zn from ZnSO<sub>4</sub>. The filtrate (ZnSO<sub>4</sub>) is then placed in evaporating dish to evaporate most of water then it's cooled after ZnSO<sub>4</sub> crystals are formed. The crystals then filtered and squeezed between filter papers to dry.

**By Reacting Insoluble Base with Acid**

E.g. Reacting MgO with Acids



The same method as reaction of acid with metal is used, so refer to diagram and above explanation, substituting reactants and products.

**By Reacting Carbonate with Acid**

E.g. Reacting CaCO<sub>3</sub> with Acids



The same process is used as reaction of acid with metal, just that carbon dioxide is produced. Carbon dioxide can be tested by bubbling it into limewater which will turn limewater colourless to milky.

**7.4 Properties and Uses of Ammonia**

**Ammonia and its Uses**

Are produced from nitrogen reacted with hydrogen

For producing: fertilisers, nitric acid, nylon, dyes, cleaners and dry cell

**The Manufacture of Ammonia: The Haber Process**

**The Process**

Nitrogen and hydrogen are mixed together in ratio 1:3, where nitrogen is obtained from air and hydrogen is obtained from natural gas, and passed over iron-catalyst.



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**Liying Dong**



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