

## Chapter 12 Study Guide - Stoichiometry

TEST: Thursday 3/10/11

### 12.1 The Arithmetic of Equations

1. Define *stoichiometry*
  
  
  
  
  
  
  
  
  
  
2. What is always conserved in every chemical reaction?

### 12.2 Chemical Calculations

1. Define *mole ratio*
  
  
  
  
  
  
  
  
  
  
2. In a balanced chemical equation, where can one find the numbers to make up the mole ratio?
  
  
  
  
  
  
  
  
  
  
3. For the following reaction:  $2\text{H}_2\text{S (g)} + 3\text{O}_2 \text{(g)} \rightarrow 2\text{SO}_2 \text{(g)} + 2\text{H}_2\text{O (g)}$ 
  - a. What is the mole ratio of hydrogen sulfide to water?
  
  
  
  
  
  
  
  
  
  
  - b. What is the mole ratio of oxygen to sulfur dioxide
  
  
  
  
  
  
  
  
  
  
  - c. What is the mole ratio of water to oxygen?

# Chapter 12 Study Guide For Content Mastery Stoichiometry

**Michael Hoelscher**



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