



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 5.1 MOLECULAR COMPOUNDS

- Classify each of the following as either an ionic or a molecular compound.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - C_2H_6
 - H_2SO_4
 - CaCl_2
 - $\text{C}_2\text{H}_5\text{OH}$
 - CO_2
 - H_2O_2
 - CH_4
 - H_2S

- Classify each of the following as either an ionic or a molecular compound.
 - K_2SO_4
 - C_2H_6
 - H_2O
 - HCl
 - CaCl_2
 - $\text{C}_2\text{H}_5\text{OH}$
 - CO_2
 - H_2O_2
 - CH_4
 - H_2S
- Write the formula for each of the following molecular compounds.
 - Carbon dioxide
 - Hydrogen peroxide
 - Hydrogen cyanide
 - Carbon monoxide
 - Diphosphorus pentoxide
 - Diphosphorus trioxide
 - Diphosphorus pentasulfide
 - Diphosphorus trisulfide
 - Diphosphorus tetrasulfide
 - Diphosphorus hexasulfide
 - Diphosphorus heptasulfide
 - Diphosphorus octasulfide
 - Diphosphorus nonasulfide
 - Diphosphorus decasulfide
 - Diphosphorus undecasulfide
 - Diphosphorus dodecasulfide
 - Diphosphorus trisulfide
 - Diphosphorus tetrasulfide
 - Diphosphorus pentasulfide
 - Diphosphorus hexasulfide
 - Diphosphorus heptasulfide
 - Diphosphorus octasulfide
 - Diphosphorus nonasulfide
 - Diphosphorus decasulfide
 - Diphosphorus undecasulfide
 - Diphosphorus dodecasulfide

SECTION 5.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride. $\text{H}-\text{F}$
- Show covalent bonding in methane. C with four H atoms around it.
- Show covalent bonding in carbon dioxide. C with two O atoms around it.
- Show covalent bonding in ethane. $\text{H}-\text{C}-\text{C}-\text{H}$ with appropriate H atoms.
- Show covalent bonding in acetylene. $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw Lewis structures for the following molecules.
 - C_2H_6
 - C_2H_4
 - C_2H_2
 - $\text{C}_2\text{H}_6\text{O}$
 - $\text{C}_2\text{H}_6\text{O}_2$
 - $\text{C}_2\text{H}_6\text{O}_3$
 - $\text{C}_2\text{H}_6\text{O}_4$
 - $\text{C}_2\text{H}_6\text{O}_5$
 - $\text{C}_2\text{H}_6\text{O}_6$
 - $\text{C}_2\text{H}_6\text{O}_7$
 - $\text{C}_2\text{H}_6\text{O}_8$
 - $\text{C}_2\text{H}_6\text{O}_9$
 - $\text{C}_2\text{H}_6\text{O}_{10}$
 - $\text{C}_2\text{H}_6\text{O}_{11}$
 - $\text{C}_2\text{H}_6\text{O}_{12}$
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 - $\text{C}_2\text{H}_6\text{O}_{99}$
 - $\text{C}_2\text{H}_6\text{O}_{100}$

SECTION 5.3 BONDING: THE ORBITALS

- Write the formula for each of the following molecular compounds.
 - Carbon dioxide
 - Hydrogen peroxide
 - Hydrogen cyanide
 - Carbon monoxide
 - Diphosphorus pentoxide
 - Diphosphorus trioxide
 - Diphosphorus pentasulfide
 - Diphosphorus trisulfide
 - Diphosphorus tetrasulfide
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Mark Steyvers



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