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## MASS PLUMBING CODE QUESTIONS AND ANSWERS

How many foot of head do you need for a test on the sanitary drainage system -  
Answers- 10'

The min spacing for hangers on horizontal 1" copper - Answers- 6'

No vent shall be installed within \_\_\_\_\_ of the trap weir - Answers- 2 pipe diameters

When sheet copper is used in plumbing it shall have a min thickness of - Answers- 12  
oz per sq ft

The storage time of a sump shall not exceed - Answers- 12 hours

Min drain size on residential laundry - Answers- 1 1/2"

Duplex pumps are required on more than \_\_\_\_\_ water closets - Answers- 5

The min spacing for hangers on horizontal pex is - Answers- 32"

Max distance of 2" trap from vent - Answers- 6'

Max distance from trap to vent 1 1/2" - Answers- 5'

A waste pipe that discharges through an air gap into a properly trapped and vented  
fixture is called - Answers- Indirect waste pipe

Min spacing on horizontal 2" copper is - Answers- 10'

Max flow from a shower head - Answers- 2.5 gpm

Min trap seal - Answers- 2"

Trap size of a tub with overhead shower - Answers- 1 1/2"

The min clearance for access to a 3" or larger clean out is - Answers- 18"

How many fixture units can a 2" trap hold - Answers- 3

1 1/2" pipe is how many fixture units - Answers- 2

Clean out plugs can be made of - Answers- Brass or plastic

Copper tubing with smallest inside diameter - Answers- Type k

Max drainage fixture units on a 3" soil stack - Answers- 48

# Plumbing Code Questions And Answers

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web the molar volume will be the volume occupied by one molar mass  $32 \text{ g}$  of the liquid expressing the density in liters instead of ml we have  $v = \frac{m}{d} = \frac{32 \text{ g}}{1.790 \text{ g/l}} = 17.88 \text{ l}$   $17.88 \text{ l}$  nonnumber

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point to keep in mind is that  $n$  can be moles of any gas since one mole always contains  $6.02 \times 10^{23}$  molecules and one molecule of any gas at the same temperature will have the same kinetic energy based on this principle it is found

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web apr 22 2020 molar volume of gases 1 counting in moles one mole is defined as  $6.02 \times 10^{23}$  particles this is the avogadro number an unchanging constant whether you are counting hydrogen atoms carbon dioxide molecules or carbonate ions a mole is always  $6.02 \times 10^{23}$  but as different particles have different mass their molar mass varies

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web the molar volume has the si unit of cubic metres per mole  $\text{m}^3 \text{ mol}^{-1}$  although it is more typical to use the units cubic decimetres per mole  $\text{dm}^3 \text{ mol}^{-1}$  for gases and cubic centimetres per mole  $\text{cm}^3 \text{ mol}^{-1}$  for liquids and solids

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web zero degrees celsius or 273 k is defined as standard temperature and one atmosphere or 760 mm hg is defined as standard pressure the volume occupied by one mole of a gas is called molar volume therefore the molar volume of a gas at standard temperature and pressure stp is 22.4 l

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web jun 8 2023  $v = 1.00 \text{ mol} \cdot 0.08206 \text{ latm mol} \cdot 273 \text{ k} = 1.00 \text{ atm} \cdot v = 22.4 \text{ l}$  in other words 1 mole of a gas will occupy 22.4 l at stp assuming ideal gas behavior at stp the volume of a gas is only dependent on number of moles of

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