

1-6, 8, 9, 11

Hydrolysis of Salts and pH of Buffer Solutions

Pre-lab Questions

Before beginning this experiment in the laboratory, you should be able to answer the following questions.

1. Define Brønsted-Lowry acids and bases.

According to Brønsted-Lowry definitions, an acid is a proton donor and a base is a proton acceptor.

2

- Which of the following ions will react with water in a hydrolysis reaction:

K^+ , Ba^{2+} , Cu^{2+} , Zn^{2+} , F^- , SO_3^{2-} , and Cl^- ?

Cu^{2+} , Zn^{2+} , F^- and SO_3^{2-} will undergo hydrolysis.

3

- For those ions in question 2 that undergo hydrolysis, write net ionic equations for the hydrolysis reaction.

Example: $Cu^{2+}(aq) + 2H_2O(l) \rightleftharpoons Cu(OH)_2(s) + 2H^+(aq)$.

4

- The K_a for HCN is 4.9×10^{-10} . What is the value of K_b for CN^- ?

$$K_b = \frac{1.0 \times 10^{-14}}{4.9 \times 10^{-10}} = 2.0 \times 10^{-5}$$

5. What are the conjugate base and conjugate acid of $H_2PO_4^-$?

The conjugate base is HPO_4^{2-} and the conjugate acid is H_3PO_4 .

6

- From what acid and what base were the following salts made: K_2SO_4 , NH_4Cl , and $Ca(NO_3)_2$?

For Ca^{2+} and SO_4^{2-} salt is made from $Ca(OH)_2$ and H_2SO_4 .

7

- Define the term *salt*.

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Michael Brown



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1. Understanding the eBook Ph Properties Of Buffer Solutions Answer Key Pre Lab
 - The Rise of Digital Reading Ph Properties Of Buffer Solutions Answer Key Pre Lab
 - Advantages of eBooks Over Traditional Books
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 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ph Properties Of Buffer Solutions Answer Key Pre Lab
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ph Properties Of Buffer Solutions Answer Key Pre Lab
 - Personalized Recommendations

- Ph Properties Of Buffer Solutions Answer Key Pre Lab User Reviews and Ratings
- Ph Properties Of Buffer Solutions Answer Key Pre Lab and Bestseller Lists
- 5. Accessing Ph Properties Of Buffer Solutions Answer Key Pre Lab Free and Paid eBooks
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 - Ph Properties Of Buffer Solutions Answer Key Pre Lab Budget-Friendly Options
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 - Highlighting and Note-Taking Ph Properties Of Buffer Solutions Answer Key Pre Lab
 - Interactive Elements Ph Properties Of Buffer Solutions Answer Key Pre Lab
- 8. Staying Engaged with Ph Properties Of Buffer Solutions Answer Key Pre Lab
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ph Properties Of Buffer Solutions Answer Key Pre Lab
- 9. Balancing eBooks and Physical Books Ph Properties Of Buffer Solutions Answer Key Pre Lab
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ph Properties Of Buffer Solutions Answer Key Pre Lab
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ph Properties Of Buffer Solutions Answer Key Pre Lab
 - Setting Reading Goals Ph Properties Of Buffer Solutions Answer Key Pre Lab
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ph Properties Of Buffer Solutions Answer Key Pre Lab
 - Fact-Checking eBook Content of Ph Properties Of Buffer Solutions Answer Key Pre Lab

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
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