

Determination of Aspirin using Back Titration

This experiment is designed to illustrate techniques used in a typical *indirect* or *back* titration. You will use the NaOH you standardized last week to back titrate an aspirin solution and determine the concentration of aspirin in a typical analgesic tablet. You will be graded on your accuracy.

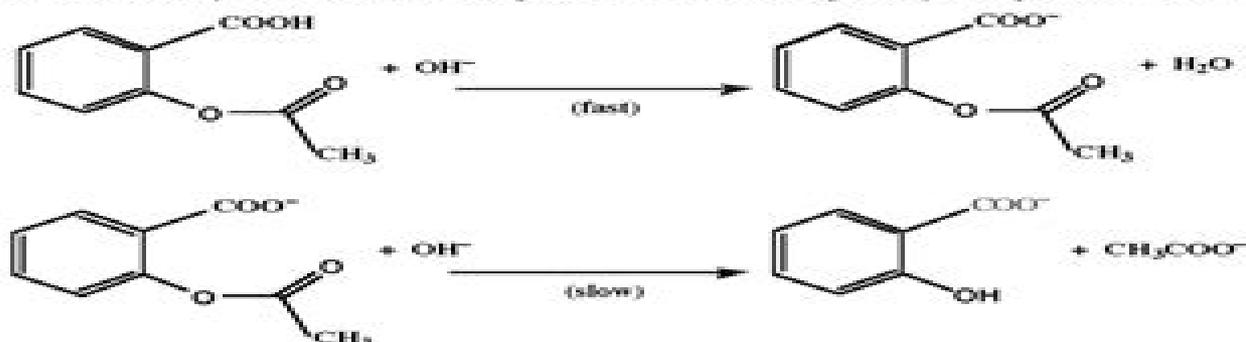
Required Reading

D.C. Harris, *Quantitative Chemical Analysis* (6th ed., W. H. Freeman, NY, 2003) p. 130, 15.

PreLab Quiz Topics

In addition to being able to explain the purpose of your experiment, the general procedure steps, the use of all chemicals in this experiment and any specific hazards, your prelab quiz may include explanations of any of the following terms: back titration, percent composition, the transition range of an indicator. You should be able to discuss the difference between a direct and a back titration, and know all relevant chemical reactions for this experiment. Describe the chemical and/or physical processes that will occur when you reach the endpoint of your titration. Explain clearly why we use a back titration for this particular analysis; i.e., why can't we just use the endpoint found from the initial NaOH titration?

Many reactions are slow or present unfavorable equilibria for direct titration. Aspirin is a weak acid that also undergoes slow hydrolysis; i.e., each aspirin molecule reacts with two hydroxide ions. To overcome this problem, a known **excess** amount of base is added to the sample solution and an HCl titration is carried out to determine the amount of unreacted base. This is subtracted from the initial amount of base to find the amount of base that actually reacted with the aspirin and hence the quantity of aspirin in the analyte.



Chemicals and their Location

Balance Room

Aspirin tablets (**Do not** bring your own sample)

Above/on lab bench

Ethanol

Phenolphthalein, indicator soln

In the Hood

Hydrochloric acid, conc., 37 wt %

Your Drawer

Sodium hydroxide, standardized soln

Analysis Of Aspirin Lab Report Conclusion

Lingsheng Yao



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