

Name _____ Date _____ Class _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then use data in the table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of O_2 , H_2 , and H_2O are 32.00 g/mol, 2.016 g/mol, and 18.02 g/mol, respectively.

Amount of O_2	Amount of H_2	Amount of H_2O	Limiting Reactant	Amount of Mass of Excess Reactant
1 molecule	1 molecule	1 molecule	none	none
2 molecules	2 molecules	2 molecules	H_2	1 molecule O_2
3 molecules	3 molecules	3 molecules	O_2	1 molecule H_2
1.00 mol	2.00 mol	1.00 mol	H_2	1.00 mol O_2
2.00 mol	4.00 mol	2.00 mol	H_2	2.00 mol O_2
3.00 mol	6.00 mol	3.00 mol	O_2	3.00 mol H_2
1.00 mol	2.00 mol	1.00 mol	H_2	1.00 mol O_2
2.00 mol	4.00 mol	2.00 mol	H_2	2.00 mol O_2
3.00 mol	6.00 mol	3.00 mol	O_2	3.00 mol H_2
32.00 g	4.032 g	18.02 g	H_2	32.00 g O_2
64.00 g	8.064 g	36.04 g	H_2	64.00 g O_2
96.00 g	12.096 g	54.06 g	O_2	96.00 g H_2

balanced equation mole ratio = 2 mol H_2 /1 mol O_2

32.00 g O_2 = 1 mol O_2 /32.00 g O_2 = 0.3125 mol O_2

20.00 g H_2 = 1 mol H_2 /2.016 g H_2 = 9.920 mol H_2

actual mole ratio = 9.920 mol H_2 /0.3125 mol O_2 = 3.176 mol H_2 /1.000 mol O_2

Because the actual mole ratio of H_2 / O_2 is larger than the balanced equation mole ratio of H_2 / O_2 , there is an excess of H_2 ; O_2 is the limiting reactant.

Mass of H_2 used = 0.3125 mol O_2 \times 2 mol H_2 /1 mol O_2 = 0.6250 mol H_2

0.6250 mol H_2 \times 2.016 g H_2 /1 mol H_2 = 1.260 g H_2

Mass of H_2 produced = 0.6250 mol H_2 \times 2.016 g H_2 /1 mol H_2 = 1.260 g H_2

Excess H_2 = 20.00 g H_2 - 1.260 g H_2 = 18.74 g H_2

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CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.4 Percent Yield

In your textbook, read about the yield of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of gallium nitride, a substance used in the manufacturing of some semiconductors for lasers.



In one experiment, the reaction yielded 7.42 g of the oxide from a 1.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_3N_2 are 69.72 g/mol and 177.00 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield volume = 7.42 g Ga_3N_2
- mass of reactant = 1.00 g Ga
- number of moles of reactant = 1.00 g Ga \times $\frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}}$ = 7.00 g Ga = 1 mol Ga/69.72 g Ga
- number of moles of product = 0.00 mol Ga \times $\frac{2 \text{ mol Ga}_3\text{N}_2}{1 \text{ mol Ga}}$ = $\frac{0.100 \text{ mol Ga}}{2 \text{ mol Ga}_3\text{N}_2}$ = 0.050 mol Ga_3N_2
- theoretical yield = 0.0500 mol Ga_3N_2 \times $\frac{177.00 \text{ g Ga}_3\text{N}_2}{1 \text{ mol Ga}_3\text{N}_2}$ = correct
- percent yield = $\frac{7.42 \text{ g Ga}_3\text{N}_2}{177.00 \text{ g Ga}_3\text{N}_2} \times 100$ = $\frac{7.42 \text{ g Ga}_3\text{N}_2}{177 \text{ g Ga}_3\text{N}_2} \times 100$

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1. Understanding the eBook Study Guide For Content Mastery Answer Key Chapter 12
 - The Rise of Digital Reading Study Guide For Content Mastery Answer Key Chapter 12
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Content Mastery Answer Key Chapter 12
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide For Content Mastery Answer Key Chapter 12
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide For Content Mastery Answer Key Chapter 12
 - Personalized Recommendations

- Study Guide For Content Mastery Answer Key Chapter 12 User Reviews and Ratings
- Study Guide For Content Mastery Answer Key Chapter 12 and Bestseller Lists
- 5. Accessing Study Guide For Content Mastery Answer Key Chapter 12 Free and Paid eBooks
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 - Interactive Elements Study Guide For Content Mastery Answer Key Chapter 12
- 8. Staying Engaged with Study Guide For Content Mastery Answer Key Chapter 12
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide For Content Mastery Answer Key Chapter 12
- 9. Balancing eBooks and Physical Books Study Guide For Content Mastery Answer Key Chapter 12
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide For Content Mastery Answer Key Chapter 12
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide For Content Mastery Answer Key Chapter 12
 - Setting Reading Goals Study Guide For Content Mastery Answer Key Chapter 12
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide For Content Mastery Answer Key Chapter 12
 - Fact-Checking eBook Content of Study Guide For Content Mastery Answer Key Chapter 12

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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