

1-45 The pressure in a vacuum chamber is measured by a vacuum gage. The absolute pressure in the chamber is to be determined.

Analysis The absolute pressure in the chamber is determined from

$$P_{abs} = P_{atm} - P_{vac} = 92 - 35 = 57 \text{ kPa}$$



1-47E The pressure in a tank in SI unit is given. The tank's pressure in various English units are to be determined.

Analysis Using appropriate conversion factors, we obtain

$$(a) \quad P = (1500 \text{ kPa}) \left(\frac{20.885 \text{ lbf/ft}^2}{1 \text{ kPa}} \right) = 31,328 \text{ lbf/ft}^2$$

$$(b) \quad P = (1500 \text{ kPa}) \left(\frac{20.885 \text{ lbf/ft}^2}{1 \text{ kPa}} \right) \left(\frac{1 \text{ ft}^2}{144 \text{ in}^2} \right) \left(\frac{1 \text{ psia}}{1 \text{ lbf/in}^2} \right) = 217.6 \text{ psia}$$

1-56 The pressure in a pressurized water tank is measured by a multi-fluid manometer. The gage pressure of air in the tank is to be determined.

Assumptions The air pressure in the tank is uniform (i.e., its variation with elevation is negligible due to its low density), and thus we can determine the pressure at the air-water interface.

Properties The densities of mercury, water, and oil are given to be 13,600, 1000, and 850 kg/m³, respectively.

Analysis Starting with the pressure at point 1 at the air-water interface, and moving along the tube by adding (as we go down) or subtracting (as we go up) the ρgh terms until we reach point 2, and setting the result equal to P_{atm} since the tube is open to the atmosphere gives

$$P_1 - \rho_{\text{water}}gh_1 + \rho_{\text{oil}}gh_2 - \rho_{\text{mercury}}gh_3 = P_{\text{atm}}$$

Solving for P_1 ,

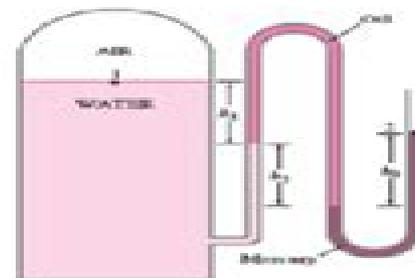
$$P_1 = P_{\text{atm}} + \rho_{\text{water}}gh_1 - \rho_{\text{oil}}gh_2 + \rho_{\text{mercury}}gh_3$$

or,

$$P_1 = P_{\text{atm}} + g(\rho_{\text{mercury}}h_3 - \rho_{\text{oil}}h_2 + \rho_{\text{water}}h_1)$$

Noting that $P_{1,gage} = P_1 - P_{\text{atm}}$ and substituting,

$$\begin{aligned} P_{1,gage} &= (9.81 \text{ m/s}^2) [(13,600 \text{ kg/m}^3)(0.4 \text{ m}) - (850 \text{ kg/m}^3)(0.2 \text{ m}) \\ &\quad + (1000 \text{ kg/m}^3)(0.3 \text{ m})] \left(\frac{1 \text{ N}}{10 \text{ kg} \cdot \text{m/s}^2} \right) \left(\frac{1 \text{ kPa}}{1000 \text{ N/m}^2} \right) \\ &= 48.9 \text{ kPa} \end{aligned}$$



Discussion Note that jumping horizontally from one tube to the next and realizing that pressure remains the same in the same fluid simplifies the analysis greatly.

Chapter 1 Solutions Thermodynamics

Michael John Misovich



Chapter 1 Solutions Thermodynamics:

Handbook of Polymer Solution Thermodynamics Ronald P. Danner, Martin S. High, 2010-09-14 Created for engineers and students working with pure polymers and polymer solutions this handbook provides up to date easy to use methods to obtain specific volumes and phase equilibrium data A comprehensive database for the phase equilibria of a wide range of polymer solvent systems and PVT behavior of pure polymers are given as are accurate predictive techniques using group contributions and readily available pure component data Two computer programs on diskettes are included POLYPROG implements procedures given for prediction and correlation for specific volume of pure polymer liquids and calculation of vapor liquid equilibria VLE of polymer solutions POLYDATA provides an easy method of accessing the data contained in the many databases in the book Both disks require a computer with a math coprocessor This handbook is a valuable resource in the design and operation of many polymer processes such as polymerization devolatilization drying extrusion and heat exchange Special Details Hardcover with Disks Special offer Purchase this book along with X 131 Handbook of Diffusion and Thermal Properties of Polymers and Polymer Solutions and receive a 20 percent discount off the list or member price

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Thermodynamics of Solutions Eli Ruckenstein, Ivan L. Shulgin, 2009-06-17 This book consists of a number of papers regarding the thermodynamics and structure of multicomponent systems that we have published during the last decade Even though they involve different topics and different systems they have something in common which can be considered as the signature of the present book First these papers are concerned with difficult or very nonideal systems i e systems with very strong interactions e g hydrogen bonding between components or systems with large differences in the partial molar volumes of the components e g the aqueous solutions of proteins or systems that are far from normal conditions e g critical or near critical mixtures Second the conventional thermodynamic methods are not sufficient for the accurate treatment of these mixtures Last but not least these systems are of interest for the pharmaceutical biomedical and related industries In order to meet the thermodynamic challenges involved in these complex mixtures we employed a variety of traditional methods but also new methods such as the fluctuation theory of Kirkwood and Buff and ab initio quantum mechanical techniques The Kirkwood Buff KB theory is a rigorous formalism which is free of any of the approximations usually used in the thermodynamic treatment of multicomponent systems This theory appears to be very fruitful when applied to the above mentioned difficult systems

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A Text Book of Thermo-chemistry and Thermodynamics Otto Sackur,1917

Modeling Thermodynamic and Diffusion Properties in Concentrated Polymer Solutions Michael John Misovich,1988 *An Introduction to the principles of physical chemistry from the standpoint of modern atomistics and thermodynamics* Edward Wight Washburn,1915

Thermodynamics and Chemistry, by F. H. MacDougall Frank Henry Macdougall,1921 *A System of Physical Chemistry: Thermodynamics* William Cudmore McCullagh Lewis,1922 Fundamentals of Engineering Thermodynamics Michael J. Moran,Howard N. Shapiro,Daisie D. Boettner,Margaret B. Bailey,2018-05-14 Fundamentals of Engineering Thermodynamics sets the standard for teaching students how to be effective problem solvers Integrated throughout this product are real world applications that emphasize the relevance of thermodynamic principles to some of the most critical problems and issues of today These include a wealth of coverage on topics related to energy and the environment biomedical bioengineering and emerging technologies The Enhanced E Text is also available bundled with an abridged print companion and can be ordered by contacting customer service here ISBN 9781119456285 Price 130 95 Canadian Price 114 95

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