

12

STOICHIOMETRY

Practice Problems

In your notebook, solve the following problems.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

1. An apple pie needs 10 large apples, 2 cups rasp and blueberries, and 1 tablespoon of cinnamon. Write a balanced equation that fits this situation. How many apples are needed to make 20 pies?
2. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate, KClO_3 . Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
3. Using the equation from problem 2, how many moles of oxygen are produced from 14 moles of potassium chlorate?
4. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from 3.8×10^{22} molecules of oxygen? How many moles of water are produced from 22.5 moles of oxygen?

SECTION 12.2 CHEMICAL CALCULATIONS

1. Calculate the number of moles of hydrogen chloride produced from 10 moles of hydrogen.



2. Calculate the number of moles of chlorine needed to form 14 moles of iron(III) chloride.



3. Calculate the number of grams of nitrogen dioxide that are produced from 4 moles of nitric oxide.



4. Calculate the mass of oxygen produced from the decomposition of 15.0 g of potassium chlorate.



5. Calculate the mass of silver needed to react with chlorine to produce 94 g of silver chloride. (Write a balanced equation first.)

6. How many liters of carbon monoxide at STP are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



7. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of nitrogen dioxide. Assume all gases are at the same conditions of temperature and pressure.



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David Kirk



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Table of Contents Chapter 12 Stoichiometry Answer

1. Understanding the eBook Chapter 12 Stoichiometry Answer
 - The Rise of Digital Reading Chapter 12 Stoichiometry Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 12 Stoichiometry Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 12 Stoichiometry Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 12 Stoichiometry Answer
 - Personalized Recommendations
 - Chapter 12 Stoichiometry Answer User Reviews and Ratings
 - Chapter 12 Stoichiometry Answer and Bestseller Lists
5. Accessing Chapter 12 Stoichiometry Answer Free and Paid eBooks

- Chapter 12 Stoichiometry Answer Public Domain eBooks
 - Chapter 12 Stoichiometry Answer eBook Subscription Services
 - Chapter 12 Stoichiometry Answer Budget-Friendly Options
6. Navigating Chapter 12 Stoichiometry Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 12 Stoichiometry Answer Compatibility with Devices
 - Chapter 12 Stoichiometry Answer Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 12 Stoichiometry Answer
 - Highlighting and Note-Taking Chapter 12 Stoichiometry Answer
 - Interactive Elements Chapter 12 Stoichiometry Answer
 8. Staying Engaged with Chapter 12 Stoichiometry Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 12 Stoichiometry Answer
 9. Balancing eBooks and Physical Books Chapter 12 Stoichiometry Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 12 Stoichiometry Answer
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chapter 12 Stoichiometry Answer
 - Setting Reading Goals Chapter 12 Stoichiometry Answer
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chapter 12 Stoichiometry Answer
 - Fact-Checking eBook Content of Chapter 12 Stoichiometry Answer
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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