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## **CHAPTER 11 INTRODUCTION TO ORGANIC CHEMISTRY**

- 11.1** Carbon can form more compounds than most other elements because carbon atoms are able to form not only single, double, and triple carbon-carbon bonds but also to link up with each other in chains and ring structures.
- 11.2** Aliphatic hydrocarbons do not contain the benzene group, or the benzene ring, whereas aromatic hydrocarbons contain one or more benzene rings.
- 11.3** Alkanes are known as saturated hydrocarbons because they contain the maximum number of hydrogen atoms that can bond with the number of carbon atoms present. Unsaturated hydrocarbons are compounds with double or triple carbon-carbon bonds. Ethane ( $\text{CH}_3\text{CH}_3$ ) is a saturated hydrocarbon, and ethylene ( $\text{CH}_2\text{CH}_2$ ) is an unsaturated hydrocarbon.
- 11.4** Isomers that differ in the order in which atoms are connected are called structural isomers.
- 11.5** Conformations are different spatial arrangements of a molecule that are generated by rotation about single bonds. In the staggered conformation of ethane, the three H atoms on one C atom are pointing away from the three H atoms on the other C atom, whereas in the eclipsed conformation the two groups of H atoms are aligned parallel to one another. A simpler and effective way of viewing these two conformations is using the

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### **Table of Contents Chang General Chemistry 7th Edition**

1. Understanding the eBook Chang General Chemistry 7th Edition
  - The Rise of Digital Reading Chang General Chemistry 7th Edition
  - Advantages of eBooks Over Traditional Books
2. Identifying Chang General Chemistry 7th Edition
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chang General Chemistry 7th Edition
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chang General Chemistry 7th Edition

- 
- Personalized Recommendations
  - Chang General Chemistry 7th Edition User Reviews and Ratings
  - Chang General Chemistry 7th Edition and Bestseller Lists
5. Accessing Chang General Chemistry 7th Edition Free and Paid eBooks
    - Chang General Chemistry 7th Edition Public Domain eBooks
    - Chang General Chemistry 7th Edition eBook Subscription Services
    - Chang General Chemistry 7th Edition Budget-Friendly Options
  6. Navigating Chang General Chemistry 7th Edition eBook Formats
    - ePub, PDF, MOBI, and More
    - Chang General Chemistry 7th Edition Compatibility with Devices
    - Chang General Chemistry 7th Edition Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chang General Chemistry 7th Edition
    - Highlighting and Note-Taking Chang General Chemistry 7th Edition
    - Interactive Elements Chang General Chemistry 7th Edition
  8. Staying Engaged with Chang General Chemistry 7th Edition
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chang General Chemistry 7th Edition
  9. Balancing eBooks and Physical Books Chang General Chemistry 7th Edition
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chang General Chemistry 7th Edition
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chang General Chemistry 7th Edition
    - Setting Reading Goals Chang General Chemistry 7th Edition
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chang General Chemistry 7th Edition

- 
- Fact-Checking eBook Content of Chang General Chemistry 7th Edition
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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