

NAME _____ DATE _____ CLASS _____

CHAPTER 12 **STUDY GUIDE FOR CONTENT MASTERY**

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and its chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27 in the space below the table.



The molar masses of O_2 , N_2 , and N_2O are 32.00 g/mol, 28.02 g/mol, and 44.02 g/mol, respectively.

Amount of O_2	Amount of N_2	Amount of N_2O	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
2 molecules	4 molecules	4 molecules	N_2	2 molecules O_2
2 molecules	4 molecules	4 molecules	2. O_2	3. 4 molecules NO
1.00 mol	2.00 mol	4. 2.00 mol	5. none	6. none
4.00 mol	8.00 mol	7. 4.00 mol	8. NO	9. 2.00 mol O_2
0.05 mol	0.10 mol	10. 0.10 mol	11. NO	12. 1.50 mol O_2
1.00 mol	2.00 mol	13. 2.00 mol	14. O_2	15. 2.00 mol NO
0.200 mol	0.200 mol	16. 0.200 mol	17. NO	18. 0.400 mol O_2
32.00 g	56.04 g	19. 92.02 g	20. none	21. none
64.00 g	112.08 g	22. 44.02 g	23. O_2	24. 50.12 g NO
64.00 g	112.08 g	25. 28.76 g	26. O_2	27. 1.24 g NO

balanced equation mole ratio = 2 mol NO /1 mol O_2

16.00 g O_2 = 1 mol O_2 /32.00 g O_2 = 0.5125 mol O_2

20.00 g NO = 1 mol NO /30.01 g NO = 0.6664 mol NO

actual mole ratio = 0.6664 mol NO /0.5125 mol O_2 = 2.132 mol NO /1.000 mol O_2

Because the actual mole ratio of NO / O_2 is larger than the balanced equation mole ratio of NO / O_2 , there is an excess of NO ; O_2 is the limiting reactant.

Mass of NO used = 0.5125 mol O_2 = 2 mol NO /1 mol O_2 = 0.4250 mol NO

0.4250 mol NO = 30.01 g NO /1 mol NO = 12.76 g NO

Mass of NO_2 produced = 0.4250 mol NO_2 = 44.02 g NO_2 /1 mol NO_2 = 18.76 g NO_2

Excess NO = 20.00 g NO = 12.76 g NO = 7.24 g NO

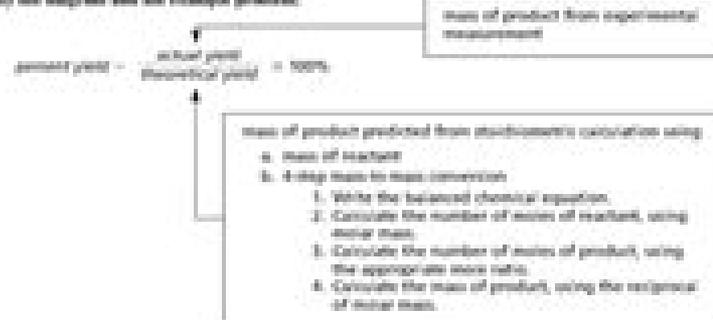
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CHAPTER 12 **STUDY GUIDE FOR CONTENT MASTERY**

Section 12.4 Percent Yield

In your textbook, read about the percent of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of yellow oxide, a substance used in the manufacturing of some semiconductor devices.



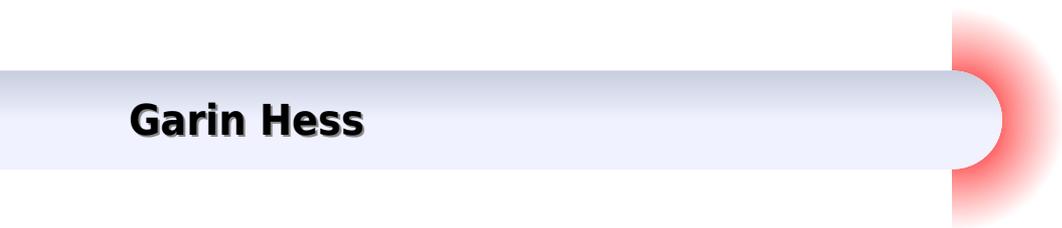
In one experiment, the reaction yielded 7.42 g of the oxide from a 7.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 187.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield: unknown 7.42 g Ga_2O_3
- mass of reactant: 7.00 g Ga correct
- number of moles of reactant: 7.00 g Ga = $\frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}}$ 7.00 g Ga = 1 mol Ga/69.72 g Ga
- number of moles of product: 0.100 mol Ga = $\frac{2 \text{ mol Ga}_2\text{O}_3}{1 \text{ mol Ga}}$ 0.100 mol Ga = 2 mol Ga_2O_3 /4 mol Ga
- theoretical yield: 0.200 mol Ga_2O_3 = $\frac{187.44 \text{ g Ga}_2\text{O}_3}{1 \text{ mol Ga}_2\text{O}_3}$ correct
- percent yield: $\frac{7.42 \text{ g Ga}_2\text{O}_3}{187.44 \text{ g Ga}_2\text{O}_3} \times 100$ 7.42 g Ga_2O_3 /187.44 g Ga_2O_3 = 100

Standardized Test Practice Chapter 12 Chemistry Answers

Garin Hess



Standardized Test Practice Chapter 12 Chemistry Answers:

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Fundamentals Of Aquatic Toxicology Gary M. Rand,2020-08-06 This text is divided into three parts The first part describes basic toxicological concepts and methodologies used in aquatic toxicity testing including the philosophies underlying testing strategies now required to meet and support regulatory standards The second part of the book discusses various factors that affect transport transformation ultimate distribution and accumulation of chemicals in the aquatic environment along with the use of modelling to predict fate The final section of the book reviews types of effects or endpoints evaluated in field studies and the use of structure activity relationships in aquatic toxicology to predict biological activity and physio chemical properties of a chemical This section also contains an extensive background of environmental legislation in the USA and within the European Community and an introduction to hazard risk assessment with case studies

How to Prepare for the SAT II Joseph A. Mascetta,1998 An extensive review of chemistry topics a full length diagnostic test precedes four up to date full length model SAT II Chemistry Tests All tests are followed with answers charts students can use to diagnose their individual study needs Chemistry topics reviewed include atomic structure bonding all other topics included in the SAT II test Joseph A Mascetta is a science educational consultant a former principal chemistry teacher at Mount Lebanon High School Pittsburgh PA

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The Top Books of the Year Standardized Test Practice Chapter 12 Chemistry Answers The year 2023 has witnessed a remarkable surge in literary brilliance, with numerous engrossing novels captivating the hearts of readers worldwide. Lets delve into the realm of popular books, exploring the engaging narratives that have enthralled audiences this year.

Standardized Test Practice Chapter 12 Chemistry Answers : Colleen Hoover's "It Ends with Us" This touching tale of love, loss, and resilience has captivated readers with its raw and emotional exploration of domestic abuse. Hoover skillfully weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can succeed.

Standardized Test Practice Chapter 12 Chemistry Answers : Taylor Jenkins Reid's "The Seven Husbands of Evelyn Hugo" This intriguing historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reid's captivating storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery.

Standardized Test Practice Chapter 12 Chemistry Answers : Delia Owens' "Where the Crawdads Sing" This captivating coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens weaves a tale of resilience, survival, and the transformative power of nature, captivating readers with its evocative prose and mesmerizing setting.

These bestselling novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of captivating stories waiting to be discovered.

The novel begins with Richard Pape, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts.

The Secret History is an exceptional and suspenseful novel that will keep you wondering until the very end. The novel is a cautionary tale about the dangers of obsession and the power of evil.

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