

Standard Enthalpy of Formation (ΔH°_f) at 298 K

Formula	KJ/mol	Formula	KJ/mol	Formula	KJ/mol
Carbon		Chlorine		Sodium	
C (s, graphite)	0	Cl (g)	121.3	Na (s)	0
C (s, diamond)	1.88	Cl ₂ (g)	0	Na (g)	107.5
CH ₄ (g)	-74.6	HCl (g)	-92.3	NaCl (s)	-411.2
CH ₃ OH (l)	-238.6	Fluorine		Na ₂ CO ₃ (s)	-1130.7
C ₂ H ₂ (g)	227.4	F (g)	79.38	NaHCO ₃ (s)	-950.8
C ₂ H ₄ (g)	52.4	F ₂ (g)	0	Sulfur	
C ₂ H ₆ (g)	-84.68	HF (g)	-273.3	S ₈ (s, rhombic)	0
C ₂ H ₅ OH (l)	-277.6	Nitrogen		S ₈ (s, monoclinic)	0.3
C ₃ H ₈ (g)	-103.85	N ₂ (g)	0	SO ₂ (g)	-296.8
C ₃ H ₆ O (l, acetone)	-248.4	NH ₃ (g)	-45.9	SO ₃ (g)	-395.7
Hydrogen		NH ₄ (s)	-365.6	H ₂ SO ₄ (l)	-814.0

Standard Enthalpy Of Formation For Various Compounds

B Lingard



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