

## Acids & Bases

### Arrhenius Acids & Bases

- Arrhenius acid: produces  $H^+$
- Arrhenius base: produces  $OH^-$
- Only applies to aqueous solutions & not all bases have  $OH^-$

### Bronsted Lowry

- Acids = proton donor
- Bases = proton acceptor
- Works in non aqueous

### Protons

- proton= $H^+$ = $H_3O^+$

### Conjugate Acids & Bases

- $H^+$  transferred to  $OH^-$

### $K_a$ and $K_b$

- Acid dissociation constant
- Base dissociation constant
- $K_a = \frac{[H_3O^+][A^-]}{[HA]}$
- $K_b = \frac{[BH][OH^-]}{[B]}$

### Acid Strengths

- Strong acid  $\rightarrow K_a > 1$ 
  - Assume 100% dissociation
  - Yields weak conjugate base
  - $HCl, HBr, H_2SO_4, HClO_4, HI, HNO_3$
- Weak acid  $\rightarrow K_a < 1$ 
  - Yields strong conjugate base
- When calculating
  - Strong & strong  $\rightarrow$  add concentrations & use
  - Strong & weak  $\rightarrow$  use strong
  - Weak & weak  $\rightarrow$  compare  $K_a$  & use the stronger one
- Polyprotic acid
  - Multiple protons

### Water

- Acid & base
- Amphoteric  $\rightarrow$  behave as either
- Autoionization
  - $2 H_2O \rightleftharpoons H_3O^+ + OH^-$
- $K_w = 1.0 \times 10^{-14}$  at 25 degrees celsius
  - Neutral water  $[H_3O^+] = [OH^-] = 1.0 \times 10^{-7}$

### pH & pOH

- $pH = -\log[H_3O^+]$
- $pOH = -\log[OH^-]$
- $pOH + pH = 14$
- $[H_3O^+] = 10^{-pH}$

# Chapter 15 Review Acids Bases

**K Morrison**



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