

# Brønsted-Lowry Acids and Bases

1 Identify the hydrogen ion donor(s) and acceptor(s) in each of the following reactions:

	H <sup>+</sup> donors (the acids)	H <sup>+</sup> acceptors (the bases)
a) $\text{HNO}_3 (\text{aq.}) + \text{H}_2\text{O} (\text{l}) \rightarrow \text{H}_3\text{O}^+ (\text{aq.}) + \text{NO}_3^- (\text{aq.})$	_____	_____
b) $\text{C}_2\text{H}_5\text{NH}_2 (\text{aq.}) + \text{H}_2\text{O} (\text{l}) \rightarrow \text{C}_2\text{H}_5\text{NH}_3^+ (\text{aq.}) + \text{OH}^- (\text{aq.})$	_____	_____
c) $\text{CH}_3\text{COOH} (\text{aq.}) + \text{H}_2\text{O} (\text{l}) \rightarrow \text{CH}_3\text{COO}^- (\text{aq.}) + \text{H}_3\text{O}^+ (\text{aq.})$	_____	_____

2 Write the formulas for the conjugate base of each of the following acids.

a) $\text{H}_2\text{SO}_3$ _____	b) $\text{HCO}_3^-$ _____	c) $\text{NH}_4^+$ _____
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3 Write the formulas for the conjugate acid of each of the following bases.

a) $\text{H}_2\text{O}$ _____	b) $\text{CO}_3^{2-}$ _____	c) $\text{PH}_3$ _____
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4 For each molecule or ion in the table, identify whether it can act as an acid or a base. Put a checkmark under each theory that describes it. There can be more than one theory per compound.

Molecule/ion	Acid or Base	Arrhenius	Brønsted-Lowry	Lewis
$\text{Br}^-$				
$\text{CN}^-$				
$\text{H}_2\text{CO}_3$				
$\text{NH}_3$				
$\text{HNO}_2$				
$\text{Ba}(\text{OH})_2$				
$\text{HCl}$				
$\text{AlCl}_3$				
$\text{Cl}^-$				
$\text{KOH}$				
$\text{IO}_3^-$				
$\text{CH}_3\text{COOH}$				

# Chapter 14 Acids And Bases Worksheet

**G Psacharopoulos**



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