
CHAPTER 11

EMULSION BREAKING

An emulsion is an intensive mixture of two liquid phases, such as oil and water, in which the liquids are mutually insoluble and where either phase may be dispersed in the other. In water chemistry, two types of emulsions are commonly found, oily wastewater and emulsified oil water or O/W emulsions and water oil emulsions (water emulsified in oil or W/O emulsions).

Oily water and water oil emulsions can usually be differentiated visually. O/W emulsions appear to be just oily, dirty water; a drop of the emulsion added to water disperses (Figure 11.1a). A W/O emulsion is usually thick and viscous; a drop of this emulsion added to water does not disperse (Figure 11.1b).

OIL-IN-WATER EMULSIONS

An oily water emulsion, in which oil is dispersed in the water phase, may consist any of various types of oil in a wide range of concentrations. These oils are defined as substances that can be extracted from water by hexane, carbon tetrachloride, chloroform, or benzene. In addition to oils, typical contaminants of these emulsions may be solids, silt, silted particles, emulsifiers, floccants, soaps, surfactants, and other solids. The types of oils found in these emulsions will depend on the industry. They may be bita, lubricants, cooking fluids, heavy hydrocarbons such as tars, grease, crude oils, and diesel oils, and also light hydrocarbons including gasoline, kerosene, and jet fuel. Their concentrations in the wastewater may vary from only a few parts per million to as much as 1 to 20% by volume.

A water O/W emulsion is a colloidal system of electrically charged oil droplets surrounded by an ionic environment. Vigorous mixing and shearing of oily wastewater in transfer storage disperses these minute oil droplets throughout the water. Emulsion stability is maintained by a combination of physical and chemical mechanisms. These emulsions act similar in behavior to the colloidal systems encountered in suspensions (settling) and rivers (silt).

One such stabilizing mechanism, emulsions, is brought about by the addition of surface-active agents, such as organic materials or chemicals, which aid in maintaining a stable colloidal system. These molecules usually carry an electric charge and seek out the oil-water interface of the emulsified droplet. Here, the pronounced charges cause the emulsion to be stabilized through repulsion of the surrounding charged droplets. Neutral compounds nevertheless can also stabilize an emulsion, since these molecules are bifunctional: one end is soluble in water, and the other end is hydrocarbon, so the molecule bridges the interface and stabilizes it.

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Michal Rosen-Zvi



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