

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the structure corresponding to the term in the vocabulary list.

- | | |
|-------------------------|---------------------------|
| _____ mole | _____ percent composition |
| _____ empirical formula | _____ empirical formula |
| _____ molar mass | _____ molar mass |
| _____ molar mass | _____ molar mass |
10. _____ is a description of the relative amounts of each element in a compound. _____ percent composition
11. _____ is the lowest whole number ratio of the atoms of the elements in a compound. _____ empirical formula
12. _____ is the mass of one mole of any gas at STP. _____ molar mass
13. _____ is the mass of one mole of a compound. _____ molar mass
14. _____ is the amount representing $n \times 6.02 \times 10^{23}$ representative particles of a substance. _____ mole
15. _____ is 6.02×10^{23} particles. _____ Avogadro's number
16. _____ is the mass of one mole of a substance at STP. _____ molar mass
17. _____ is the mass of one mole of a gas at STP. _____ molar mass
18. _____ is the mass of one mole of a gas at STP. _____ molar mass

Part 2: Section Concept Review

10.1 The Mole: A Fundamental

Complete the following activities to demonstrate your knowledge of the terms and your understanding of the concepts involved with the terms. Each activity has a corresponding question sheet to be given as a follow-up.

1. Drawing whole table chemistry. At STP, what is the volume for a single quantity called the _____ mole. _____ Representative particles of particles for the gas volume is _____ Avogadro's number.

2. The mass of one mole of a compound, whether gas or liquid, is _____ molar mass. _____ of the atoms in the compound. When you substitute the molar mass for the volume, you obtain the _____ molar mass _____ of the compound. These are _____ 6.02×10^{23} _____ representative particles are made of one substance.

Practice Problems

- 10.1.1. Calculate the molar mass of each of the following:
- 10.1.2. Calculate the molar mass of each of the following:
- 10.1.3. Calculate the molar mass of each of the following:
- 10.1.4. Calculate the molar mass of each of the following:

Answers

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WJ Hussar



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