Reduction : copper oxide is being reduced, CuO is an oxidising agent
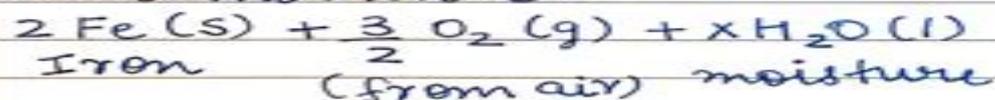
$$CuO + H_2 \xrightarrow{\text{Heat}} Cu + H_2O$$
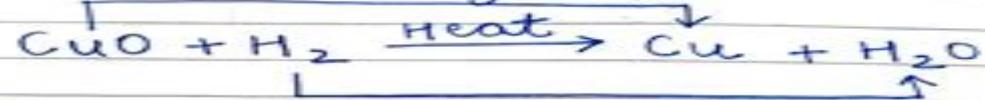
Oxidation : Hydrogen is being oxidised, H_2 is the reducing agent.

OXIDATION
↓
EFFECTS IN EVERYDAY LIFE

CORROSION
Slow conversion of metals into their undesirable compounds (sulphides, carbonates, oxides, etc.) by interaction with atmospheric gases and moisture

$2\,Fe\,(s) + \dfrac{3}{2}\,O_2\,(g) + x\,H_2O\,(l)$
Iron   (from air) moisture

$\longrightarrow FeO_3 . x\,H_2O\,(s)$
Hydrated ferric oxide (Rust)

RANCIDITY
Oxidation of oils or fats in food resulting in bad taste and smell.

METHODS OF PREVENTION

By adding Antioxidants ←
Vacuum ~~Vaccu~~ Packing ←
Replacing air by Nitrogen ←
Refrigeration of the Foodstuff ←

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