

## 17. SOLUBILITY PRODUCT CONSTANT

### GOALS

1. Become aware of solubility product constants
2. Estimate a solubility constant
3. Observe that a solubility constant is a special case of an equilibrium constant

### BACKGROUND

Ionic compounds, such as  $\text{NaCl}$ , dissociate to form ions when they dissolve in water. An example is sodium chloride:



Where  $aq$  is a solution and  $s$  is a solid.

Solid  $\text{NaCl}$  dissociates to become  $\text{Na}^+$  and  $\text{Cl}^-$  ions. The  $\text{Na}^+$  cations are attracted to the negative "ends" of water molecules, and the  $\text{Cl}^-$  anions are attracted to the positive "ends" of water molecules.

In other compounds the attraction between the cations and the anions is stronger and there is little dissociation. These are sparingly or slightly soluble salts. An example is silver chloride,  $\text{AgCl}$ . A beaker of this solution often has solid  $\text{AgCl}$  crystals in the bottom. The crystals are in equilibrium with the  $\text{Ag}^+$  and  $\text{Cl}^-$  ions that are in solution:



In this reaction the rate of  $\text{AgCl}$  dissolving is equal to the rate of  $\text{Ag}^+$  and  $\text{Cl}^-$  ions combining to form solid  $\text{AgCl}$ . In other words, the reaction is going both ways at the same time and at the same rate. It is said to be in equilibrium. The equilibrium constant is:



# Solubility Product Constant Lab Activity

**Rachel S Tattersall**



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## **Table of Contents Solubility Product Constant Lab Activity**

1. Understanding the eBook Solubility Product Constant Lab Activity
  - The Rise of Digital Reading Solubility Product Constant Lab Activity
  - Advantages of eBooks Over Traditional Books
2. Identifying Solubility Product Constant Lab Activity
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Solubility Product Constant Lab Activity
  - User-Friendly Interface
4. Exploring eBook Recommendations from Solubility Product Constant Lab Activity
  - Personalized Recommendations
  - Solubility Product Constant Lab Activity User Reviews and Ratings
  - Solubility Product Constant Lab Activity and Bestseller Lists

5. Accessing Solubility Product Constant Lab Activity Free and Paid eBooks
  - Solubility Product Constant Lab Activity Public Domain eBooks
  - Solubility Product Constant Lab Activity eBook Subscription Services
  - Solubility Product Constant Lab Activity Budget-Friendly Options
6. Navigating Solubility Product Constant Lab Activity eBook Formats
  - ePub, PDF, MOBI, and More
  - Solubility Product Constant Lab Activity Compatibility with Devices
  - Solubility Product Constant Lab Activity Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Solubility Product Constant Lab Activity
  - Highlighting and Note-Taking Solubility Product Constant Lab Activity
  - Interactive Elements Solubility Product Constant Lab Activity
8. Staying Engaged with Solubility Product Constant Lab Activity
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Solubility Product Constant Lab Activity
9. Balancing eBooks and Physical Books Solubility Product Constant Lab Activity
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Solubility Product Constant Lab Activity
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Solubility Product Constant Lab Activity
  - Setting Reading Goals Solubility Product Constant Lab Activity
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Solubility Product Constant Lab Activity
  - Fact-Checking eBook Content of Solubility Product Constant Lab Activity
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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