

Midterm Lab Quiz 2

1. The number of electron shells in elements of the same period is the same. **True**
2. A positively charged ion is referred to as a **cation**
3. Chemical knowledge extended throughout cultures, and by the mid-eighteenth century, 43 elements had been identified. **False**
4. **Is a form of chemical bond in which one atom loses its valence electrons and gains them from another. Ionic bond**
5. A process that transfers heat into its surroundings is known as an [1] reaction. **Exothermic**
6. The periodic table is a systematic organization of chemical elements ordered according to their electron configurations. **True**
7. The ability of an atom to attract electrons to itself is known as **electronegativity**
8. The binding strength of an ionic substance is measured by **ionic charge**
9. A negatively charged ion is known as an **anion**
10. The transfer of an electron from a metal atom to a [1] atom forms an ionic connection. **non-metal**
11. Mendeleev's periodic chart has been updated to include a total of 128 elements. The atomic number is the number above the symbol in the periodic table that represents the number of protons in the nucleus. The sum of neutrons and protons determines the atomic mass. **False**
12. Joseph Proust and others experimented with the Law of Definite Proportions at the turn of the nineteenth century. **True**
13. The origins of atomic theory may be traced back to ancient Greek and Hellenistic Egyptian intellectuals. **True**
14. The modern hunt for a universal remedy and the philosopher's stone is known as alchemy. Chemistry emerged as a result of the process. **False**
15. An [1] is an ionic chemical that dissolves in water and allows the solution to conduct electricity. **Electrolyte**
16. The nucleus of the most common isotope of hydrogen, made of two up quarks and a down quark, is a negatively charged subatomic particle that is part of the nucleus of an atom and determines the atomic number of an element. **False**
17. **is a monatomic element that is inert or non-reactive, with very few exceptions, and belongs to group 18 of the periodic table. Noble gases**
18. The Aufbau principle outlines the filling of orbitals and the construction of atoms with known electronic configurations in a step-by-step manner. **True**

Chemistry Midterm Lab Answers

SB Merriam



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