

Intro to Chemistry Final Exam Study Guide

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- _____ 1. The Fahrenheit temperature scale is defined by which of the following temperatures?
a. Water boils at 100 degrees and freezes at 32 degrees.
b. Water boils at 212 degrees and freezes at 32 degrees.
c. Water evaporates at 212 degrees and freezes at 0 degrees.
d. Liquid water turns to a gas at 100 degrees and to a solid at 0 degrees.
- _____ 2. A cold-blooded reptile basks on a warm rock in the sun. Its body is warmed by
a. radiation. c. convection.
b. conduction. d. Both (a) and (b).
- _____ 3. Temperature is a measure of the average _____ energy of an object's particles.
a. mechanical c. potential
b. kinetic d. light
- _____ 4. The temperature at which all molecular motion stops is
a. 0°C. c. 0 K.
b. 0°F. d. All of the above.
- _____ 5. Convert 500°F to degrees Celsius.
a. 260°C c. 842°C
b. 296°C d. 958°C
- _____ 6. Convert 300 K to the Celsius scale.
a. 0°C c. 73°C
b. 27°C d. 573°C
- _____ 7. Convert 458 K to the equivalent Fahrenheit temperature.
a. 185°F c. 365°F
b. 237°F d. 731°F
- _____ 8. Which of the following temperatures is impossible to measure?
a. -85°F c. -20 K
b. -50°C d. 545°F
- _____ 9. What is -175°C on the Kelvin scale?
a. 76 K c. 98 K
b. 89 K d. 448 K
- _____ 10. The principles of atomic theory recognized today were conceived by
a. Avogadro. c. Dalton.
b. Bohr. d. Rutherford.
- _____ 11. The total number of protons and neutrons in the nucleus of an atom is its
a. atomic number. c. mass number.
b. Avogadro number. d. average atomic mass.
- _____ 12. All atoms of the same element have the same
a. atomic mass. c. mass number.
b. number of neutrons. d. atomic number.
- _____ 13. Atoms of the same element can differ in
a. chemical properties. c. atomic number.
b. mass number. d. number of protons and electrons.
- _____ 14. Chlorine has atomic number 17 and mass number 35. It has
a. 17 protons, 17 electrons, and 18 neutrons.
b. 35 protons, 35 electrons, and 17 neutrons.
c. 17 protons, 17 electrons, and 52 neutrons.
d. 18 protons, 18 electrons, and 17 neutrons.
- _____ 15. Phosphorus-33 (atomic number 15) contains
a. 33 protons. c. 33 neutrons.
b. 18 neutrons. d. 18 protons.
- _____ 16. Neon-22 contains 12 neutrons. It also contains
a. 12 protons. c. 22 electrons.
b. 22 protons. d. 10 protons.
- _____ 17. Valence electrons determine an atom's

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R Barnett



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