

# AP Chemistry Final Study Guide

## Unit 1: Atomic Structure and Properties (7%–9%)

- At STP, one mole of gas always occupies 22.4 L
- **Mass Spectroscopy:**
  - Atoms have different isotopes (carbon-12, carbon-14, etc) and this will affect the average atomic mass
  - The most abundant isotope will have the greatest effect
- **Coulomb's Law:**
  - The greater the charge of the nucleus, the more energy is required to ionize an electron
  - The closer an electron is to the nucleus, the more energy is required to ionize an electron
- **Electromagnetic Radiation:**
  - When atoms absorb energy, electrons jump to higher energy levels
  - When electrons drop from higher to lower energy levels, atoms give off electromagnetic radiation energy
- **Kinetic Energy of Ionized Electrons:**
  - The faster an ejected electron is going, the more kinetic energy it has
  - Electrons originally further away from the nucleus require less energy to eject, so it will have more kinetic energy
- **Electron Configuration:**
  - When entering the d-block (transition metals), the energy level drops by 1 unless you are using shorthand notation
    - Nickel:  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^8 4s^2$  (notice the d-block)
    - Nickel (with shorthand):  $[\text{Ar}]4s^2 3d^8$
- **Law of Definite Proportions:**
  - A pure chemical compound broken down into elements always contains elements of a fixed ratio, independent of where and how it was created
- **Law of Multiple Proportions:**
  - When elements form compounds, the proportions of the elements in those chemical compounds can be expressed in small whole number ratios
- **Aufbau Principle:**
  - When building up the electron configuration of an atom, electrons are placed in orbitals, subshells, and shells in order of increasing energy

# Chemistry Final Exam Study Guide

**JS Bruner**



## **Chemistry Final Exam Study Guide:**

## Unveiling the Energy of Verbal Art: An Psychological Sojourn through **Chemistry Final Exam Study Guide**

In some sort of inundated with screens and the cacophony of quick conversation, the profound power and psychological resonance of verbal beauty frequently fade into obscurity, eclipsed by the regular onslaught of sound and distractions. Yet, nestled within the lyrical pages of **Chemistry Final Exam Study Guide**, a charming work of fictional splendor that pulses with organic emotions, lies an unforgettable trip waiting to be embarked upon. Published by way of a virtuoso wordsmith, that enchanting opus courses visitors on an emotional odyssey, gently exposing the latent possible and profound affect embedded within the complex web of language. Within the heart-wrenching expanse with this evocative evaluation, we will embark upon an introspective exploration of the book is main subjects, dissect its captivating publishing design, and immerse ourselves in the indelible effect it leaves upon the depths of readers souls.

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