

Nitrogen dioxide, NO_2 , can be made by reacting nitrogen monoxide, NO , with oxygen, O_2 .



A chemist mixes 0.450 mol NO with 0.450 mol O_2 in a container. This container is then heated and equilibrium is reached.

At equilibrium, the mixture contains 0.400 mol NO and the total pressure is 400 kPa. Calculate K_p and the units

Step 1 - Write out the equation and fill in the initial moles

| | | | | | |
|-------------|--------------|---|--------------|----------------------|----------------|
| | 2NO | + | O_2 | \rightleftharpoons | 2NO_2 |
| Initial | 0.450 | | 0.450 | | 0 |
| Change | -0.05 | | -0.025 | | +0.05 |
| Equilibrium | 0.400 | | 0.425 | | 0.05 |

Step 2 - Work out the partial pressures of each gas

$$p(\text{NO}) = 0.400 / 0.875 \times 400 = 182.86 \text{ kPa}$$

$$p(\text{O}_2) = 0.425 / 0.875 \times 400 = 194.29 \text{ kPa}$$

$$p(\text{NO}_2) = 0.05 / 0.875 \times 400 = 22.86 \text{ kPa}$$

Step 3 - Put the values in to the K_p equation and work out the answer

$$K_p = \frac{p(\text{NO}_2)^2}{p(\text{NO})^2 \times p(\text{O}_2)}$$

$$K_p = \frac{(22.86)^2}{(182.86)^2 \times (194.29)}$$

$$K_p = 8.04 \times 10^{-5}$$

Step 4 - Calculate the units

$$K_p = \frac{\cancel{(\text{kPa})^2}}{(\cancel{\text{kPa}})^2 \times (\text{kPa})}$$

$$K_p = \text{kPa}^{-1}$$

Chemistry Chapter 19 Assessment Answers

JG Myers



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