

Energy and Chemical Change

Section 16.1 Energy

In your textbook, read about the nature of energy.

In the space at the left, write *true* if the statement is true; if the statement is false, change the italicized word or phrase to make it true.

- _____ 1. *Energy* is the ability to do work or produce heat.
- _____ 2. The law of conservation of energy states that energy *can be* created and destroyed.
- _____ 3. Chemical potential energy is energy stored in a substance because of its *composition*.
- _____ 4. *Heat* is a form of energy that flows from a warmer object to a cooler object.
- _____ 5. A calorie is the amount of energy required to raise the temperature of *one gram* of pure water by one degree Celsius.
- _____ 6. A calorie is the SI unit of heat and energy.
- _____ 7. The *specific heat* of a substance is the amount of heat required to raise the temperature of *one gram* of that substance by one degree Celsius.
- _____ 8. *Kinetic energy* is energy of motion.
- _____ 9. Chemicals participating in a chemical reaction contain *only potential energy*.
- _____ 10. One nutritional Calorie is equal to *100 calories*.
- _____ 11. One calorie equals *4.184 joules*.
- _____ 12. When a fuel is burned, some of its *chemical potential energy* is lost as heat.
- _____ 13. To convert kilojoules to joules, *divide* the number of kilojoules by 1000 joules/1 kilojoule.

Answer the following question. Show all your work.

14. If the temperature of a 500.0-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J/(g·°C).

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CL Gary



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1. Understanding the eBook Chemistry Chapter 16 Study Guide For Content Mastery Answers
 - The Rise of Digital Reading Chemistry Chapter 16 Study Guide For Content Mastery Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 16 Study Guide For Content Mastery Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 16 Study Guide For Content Mastery Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 16 Study Guide For Content Mastery Answers
 - Personalized Recommendations
 - Chemistry Chapter 16 Study Guide For Content Mastery Answers User Reviews and Ratings
 - Chemistry Chapter 16 Study Guide For Content Mastery Answers and Bestseller Lists
5. Accessing Chemistry Chapter 16 Study Guide For Content Mastery Answers Free and Paid eBooks
 - Chemistry Chapter 16 Study Guide For Content Mastery Answers Public Domain eBooks
 - Chemistry Chapter 16 Study Guide For Content Mastery Answers eBook Subscription Services

- Chemistry Chapter 16 Study Guide For Content Mastery Answers Budget-Friendly Options
- 6. Navigating Chemistry Chapter 16 Study Guide For Content Mastery Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 16 Study Guide For Content Mastery Answers Compatibility with Devices
 - Chemistry Chapter 16 Study Guide For Content Mastery Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 16 Study Guide For Content Mastery Answers
 - Highlighting and Note-Taking Chemistry Chapter 16 Study Guide For Content Mastery Answers
 - Interactive Elements Chemistry Chapter 16 Study Guide For Content Mastery Answers
- 8. Staying Engaged with Chemistry Chapter 16 Study Guide For Content Mastery Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 16 Study Guide For Content Mastery Answers
- 9. Balancing eBooks and Physical Books Chemistry Chapter 16 Study Guide For Content Mastery Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 16 Study Guide For Content Mastery Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 16 Study Guide For Content Mastery Answers
 - Setting Reading Goals Chemistry Chapter 16 Study Guide For Content Mastery Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 16 Study Guide For Content Mastery Answers
 - Fact-Checking eBook Content of Chemistry Chapter 16 Study Guide For Content Mastery Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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