

## Chapter 13

# Hydrocarbons

### Question 13.1:

How do you account for the formation of ethane during chlorination of methane?

### Answer 13.1:

Chlorination of methane proceeds via a free radical chain mechanism. The whole reaction takes place in the given three steps.

#### Step 1: Initiation:

The reaction begins with the homolytic cleavage of Cl – Cl bond as:



#### Step 2: Propagation:

In the second step, chlorine free radicals attack methane molecules and break down the C–H bond to generate methyl radicals as:



These methyl radicals react with other chlorine free radicals to form methyl chloride along with the liberation of a chlorine free radical.

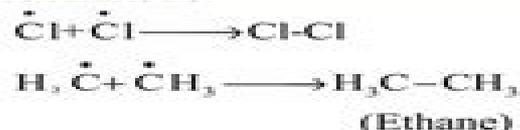


Hence, methyl free radicals and chlorine free radicals set up a chain reaction. While HCl and CH<sub>3</sub>Cl are the major products formed, other higher halogenated compounds are also formed as:



#### Step 3: Termination:

Formation of ethane is a result of the termination of chain reactions taking place as a result of the consumption of reactants as:



Hence, by this process, ethane is obtained as a by-product of chlorination of methane.

# Chemistry Chapter 13 Solutions

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## **Table of Contents Chemistry Chapter 13 Solutions**

1. Understanding the eBook Chemistry Chapter 13 Solutions
  - The Rise of Digital Reading Chemistry Chapter 13 Solutions
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 13 Solutions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Chapter 13 Solutions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 13 Solutions
  - Personalized Recommendations
  - Chemistry Chapter 13 Solutions User Reviews and Ratings
  - Chemistry Chapter 13 Solutions and Bestseller Lists
5. Accessing Chemistry Chapter 13 Solutions Free and Paid eBooks
  - Chemistry Chapter 13 Solutions Public Domain eBooks
  - Chemistry Chapter 13 Solutions eBook Subscription Services
  - Chemistry Chapter 13 Solutions Budget-Friendly Options

6. Navigating Chemistry Chapter 13 Solutions eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Chapter 13 Solutions Compatibility with Devices
  - Chemistry Chapter 13 Solutions Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Chapter 13 Solutions
  - Highlighting and Note-Taking Chemistry Chapter 13 Solutions
  - Interactive Elements Chemistry Chapter 13 Solutions
8. Staying Engaged with Chemistry Chapter 13 Solutions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Chapter 13 Solutions
9. Balancing eBooks and Physical Books Chemistry Chapter 13 Solutions
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Chapter 13 Solutions
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 13 Solutions
  - Setting Reading Goals Chemistry Chapter 13 Solutions
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 13 Solutions
  - Fact-Checking eBook Content of Chemistry Chapter 13 Solutions
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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