

KINETICS Practice Problems and Solutions

Name: _____
 Period: _____ Date: _____

AP Chemistry
 Dr. Mandes

The following questions represent potential types of quiz questions. Please answer each question completely and thoroughly. The solutions will be posted on-line on Monday.

5. Please do #18 in chapter 12 of your text.
- This increases the $[H_2]$, which will increase the rate, but has no effect on k
 - Due to Arrhenius equation, changing temperature changes the value of k .
 - Catalyst function by lowering the activation energy, so due to Arrhenius equation, changing the activation energy changes the value of k .

6. Consider the reaction: $P_4 + 6 H_2 \rightarrow 4 PH_3$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

| $[P_4]$, mol/L | $[H_2]$, mol/L | Initial Rate, mol/L · s |
|-----------------|-----------------|-------------------------|
| 0.0110 | 0.0075 | 3.20×10^{-4} |
| 0.0110 | 0.0150 | 6.40×10^{-4} |
| 0.0220 | 0.0150 | 6.39×10^{-4} |

- What is the order with respect to: P_4 0
 H_2 1
- Write the rate law for this reaction. $rate = k[H_2]$
- Determine the value and units of the rate constant, k . **plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 0.0427 s^{-1}$**

7. Consider the reaction: $SO_2 + O_3 \rightarrow SO_3 + O_2$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

| $[SO_2]$, mol/L | $[O_3]$, mol/L | Initial Rate, mol/L · s |
|------------------|-----------------|-------------------------|
| 0.25 | 0.40 | 0.118 |
| 0.25 | 0.20 | 0.118 |
| 0.75 | 0.20 | 1.062 |

- What is the order with respect to: SO_2 2
 O_3 0
- Write the rate law for this reaction. $rate = k[SO_2]^2[O_3]^0$
- Determine the value and units of the rate constant, k . **plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 2.36 molL^{-2} s^{-1}$**

8. Consider the following mechanism.
- $$A_2 + B_2 \rightarrow R + C \quad (\text{slow})$$
- $$A_2 + R \rightarrow C \quad (\text{fast})$$

- Write the overall balanced chemical equation. $2 A_2 + B_2 \rightarrow 2 C$
- Identify any intermediates within the mechanism. **R**
- What is the order with respect to each reactant? **A_2 : 1st; B_2 : 1st**

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IM Harris



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