

Chapter 2 - Practice Test #1

7. An equilibrium system shifts left when the
- rate of the forward reaction is equal to the rate of the reverse reaction.
 - rate of the forward reaction is less than the rate of the reverse reaction.
 - rate of the forward reaction is greater than the rate of the reverse reaction.
 - rate of the forward reaction and the reverse reaction are constant.
8. A 1.00 L flask contains a gaseous equilibrium system. The addition of reactants to this flask results in a
- shift left and a decrease in the concentration of the products.
 - shift left and a increase in the concentration of the products.
 - shift right and a decrease in the concentration of the products.
 - shift right and a increase in the concentration of the products.
9. Consider the following equilibrium: $\text{CH}_4\text{g} + \text{H}_2\text{Og} + \text{heat} = \text{COg} + 3\text{H}_2\text{g}$
In which of the following will both stresses shift the equilibrium to the right?
- a decrease in temperature and a decrease in volume
 - a increase in temperature and a decrease in volume
 - a decrease in temperature and a increase in volume
 - a increase in temperature and a increase in volume
10. Consider the following equilibrium: $2\text{SO}_2\text{g} + \text{O}_2\text{g} = 2\text{SO}_3\text{g}$ $\Delta H = -198 \text{ kJ}$
There will be no shift in this equilibrium when
- | | |
|--------------------------------|----------------------------------|
| A. more O_2 is added. | C. the volume is increased. |
| B. a catalyst is added. | D. the temperature is increased. |
11. Consider the following equilibrium: $2\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{Og} = \text{Fe}_2\text{O}_3\text{g} + 3\text{H}_2\text{g}$
The equilibrium expression is
- | | |
|---|--|
| A. $\text{Keq} = \frac{[\text{Fe}_2\text{O}_3][\text{H}_2]^3}{[\text{Fe}]^2[\text{H}_2\text{O}]^3}$ | B. $\text{Keq} = \frac{[\text{Fe}_2\text{O}_3][3\text{H}_2]}{[2\text{Fe}][3\text{H}_2\text{O}]}$ |
| C. $\text{Keq} = \frac{[\text{H}_2]^3}{[\text{H}_2\text{O}]^3}$ | D. $\text{Keq} = [\text{H}_2]^3$ |
12. Consider the following equilibrium: $\text{N}_2\text{O}_4\text{g} = 2\text{NO}_2\text{g}$ $\text{Keq} = 0.133$
At equilibrium, the $[\text{N}_2\text{O}_4]$ is equal to

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