



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 5.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or molecular.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - C_2H_6
 - H_2SO_4
- Write the formula for the following molecular compounds.
 - nitrogen dioxide
 - carbon dioxide
 - carbon monoxide
 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide
- Write the name for each of the following molecular compounds.
 - CO_2
 - N_2O
 - N_2O_5
 - CO
- Write the name for each of the following molecular compounds.
 - C_2H_6
 - C_2H_4
 - C_2H_2
 - $\text{C}_2\text{H}_6\text{O}$
 - $\text{C}_2\text{H}_6\text{O}_2$
 - $\text{C}_2\text{H}_4\text{O}_2$
- Write the formula for each of the following molecular compounds.
 - methane
 - ethane
 - propane
 - butane
 - pentane
 - hexane
 - heptane
 - octane
 - nonane
 - decane

SECTION 5.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride. $\text{H}-\text{F}$
- Show covalent bonding in methane. C with four H atoms bonded to it.
- Show covalent bonding in carbon dioxide. $\text{O}=\text{C}=\text{O}$
- Show covalent bonding in ethane. $\text{H}-\text{C}-\text{C}-\text{H}$ with appropriate H atoms.
- Show covalent bonding in ethyne. $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw Lewis structures for the following molecules.
 - H_2O
 - CO_2
 - C_2H_6
 - C_2H_4
 - C_2H_2
 - $\text{C}_2\text{H}_6\text{O}$
 - $\text{C}_2\text{H}_6\text{O}_2$
 - $\text{C}_2\text{H}_4\text{O}_2$

SECTION 5.3 BONDING: THE ORBITALS

- Write the hybrid orbitals for the central atom in each of the following.
 - CH_4
 - C_2H_6
 - C_2H_4
 - C_2H_2
 - $\text{C}_2\text{H}_6\text{O}$
 - $\text{C}_2\text{H}_6\text{O}_2$
 - $\text{C}_2\text{H}_4\text{O}_2$
- Write the hybrid orbitals for the central atom in each of the following.
 - CO_2
 - CO
 - NO_2
 - NO
 - SO_2
 - SO
 - O_3
 - O_2

Chapter 8 Covalent Bonding Practice Problems Answers

B Lingard



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