

Chemistry I Chapter 5 Study Guide

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- _____ 1. The noble gases, such as helium and xenon, are non-reactive because:
- they have completely filled outer electron shells.
 - they are chemically unstable.
 - they are unusually large atoms.
 - they have been around longest on the earth.
- _____ 2. Elements in the same column on the periodic table have similar chemical reactivity because:
- they contain similar amounts of energy.
 - they have the same number of valence shell electrons.
 - atoms prefer to have completely full or completely empty valence shells.
 - they have similar numbers of isotopes.
- _____ 3. Chlorine, element number 17, is located in the third row and the next-to-the-last group of the periodic table. How many electron shells would be completely filled by a neutral atom of chlorine? How many electrons would be left over?
- 1 shell filled, 7 electrons left over
 - 1 shell filled, 1 electron left over
 - 2 shells filled, 7 electrons left over
 - 2 shells filled, 1 electron left over
- _____ 4. How many energy levels are filled in a krypton atom?

36
Kr
Krypton
83.80

- Three.
 - Four.
 - Five.
 - Six.
- _____ 5. The electronegativity of an element is defined as:
- the attraction an atom has for the shared pair of electrons in a covalent bond.
 - the attraction an atom has for the shared pair of electrons in an ionic bond.
 - the amount of static electricity an element produces.
 - the amount of charge one atom of an element produces.
- _____ 6. Which of the following is true for valence electrons?
- Valence electrons are always located in the innermost energy level.
 - Valence electrons may be lost or gained by an atom when bonding occurs with another atom.
 - Each element in the periodic table has a different number of valence electrons.
 - Valence electrons are found only in radioactive isotopes.
- _____ 7. Mendeleev arranged the known chemical elements in a table according to increasing
- atomic number.
 - number of electrons.
 - number of protons.
 - mass.
- _____ 8. Which list of elements contains only metals?
- carbon, iodine, tin
 - tin, copper, cesium
 - helium, iron, copper
 - iodine, carbon, argon
- _____ 9. As you move from left to right across a period, the number of valence electrons
- increases.
 - stays the same.
 - increases and then decreases.
 - decreases.
- _____ 10. A member of the boron family has three valence electrons, while a member of the nitrogen family has
- none.
 - four.

Chapter 5 Chemistry Answers

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