

# Solutions

## Exercises

1. Define the term solution. How many types of solutions are formed? Write briefly about each type with an example.

**Ans.** A solution is a homogeneous mixture of two or more pure substances. Solute and solvent are two components of a solution.

**Types of solutions:** On the basis of different states of solute and solvent solutions are of three types mainly, which are further categorized into three.

Type of Solution	Solute	Solvent	Common Examples
Gaseous Solutions	Gas	Gas	Mixture of oxygen and nitrogen gases
	Liquid	Gas	Chloroform mixed with nitrogen gas
	Solid	Gas	Camphor in nitrogen gas
Liquid Solutions	Gas	Liquid	Oxygen dissolved in water
	Liquid	Liquid	Ethanol dissolved in water
	Solid	Liquid	Glucose dissolved in water
Solid Solutions	Gas	Solid	Solution of hydrogen in palladium
	Liquid	Solid	Amalgam of mercury with sodium
	Solid	Solid	Copper dissolved in gold

2. Give an example of a solid solution in which the solute is a gas.

**Ans.** In case a solid solution is formed between two substances, an interstitial solid solution will be formed. For example: a solution of hydrogen in palladium is a solution in which the solute is a gas.

3. Define the following terms:

(i) Mole fraction

(iii) Molarity

(ii) Molality

(iv) Mass percentage

**Ans.** (i) **Mole fraction:** It is defined as the ratio of the number of moles of the individual component (solute or solvent) to the total number of moles of all components present in the solution. If  $n_A$  is the number of moles of solute dissolved in  $n_B$  moles of solvent, then Mole fraction of solute

$$(X_A) = \frac{n_A}{n_A + n_B}$$

$$\text{Mole fraction of solvent } (X_B) = \frac{n_B}{n_A + n_B}$$

(ii) **Molality:** It is defined as the number of moles of a solute present in 1000g (1kg) of a solvent.

$$\text{Molality } (m) = \frac{\text{Number of moles of solute}}{\text{Weight of solvent in kg}} = \frac{n}{W}$$

$$\text{Unit} = \text{mol kg}^{-1} \text{ or m(molal)}$$

(iii) **Molarity:** It is defined as the number of moles of solute present in one litre of solution.

$$\text{Molarity } (M) = \frac{\text{Number of moles of solute}}{\text{Volume of solution in litre}} = \frac{n}{V}$$

$$\text{Unit} = \text{mol L}^{-1} \text{ or M(Molar)}$$

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