

- Which event must *always* occur for a chemical reaction to take place?
  - formation of a precipitate
  - formation of a gas
  - effective collisions between reacting particles
  - addition of a catalyst to the reaction system
- In most aqueous reactions as temperature increases, the effectiveness of collisions between reacting particles
  - decreases
  - increases
  - remains the same
- A 5.0-gram sample of zinc and a 50.-milliliter sample of hydrochloric acid are used in a chemical reaction. Which combination of these samples has the fastest reaction rate?
  - a zinc strip and 1.0 M HCl(aq)
  - a zinc strip and 3.0 M HCl(aq)
  - zinc powder and 1.0 M HCl(aq)
  - zinc powder and 3.0 M HCl(aq)
- A 1.0-gram piece of zinc reacts with 5 milliliters of HCl(aq). Which of these conditions of concentration and temperature would produce the greatest rate of reaction?
  - 1.0 M HCl(aq) at 20.°C
  - 1.0 M HCl(aq) at 40.°C
  - 2.0 M HCl(aq) at 20.°C
  - 2.0 M HCl(aq) at 40.°C
- At STP, which 4.0-gram zinc sample will react fastest with dilute hydrochloric acid?
  - lump
  - bar
  - powdered
  - sheet metal
- Given the balanced equation representing a reaction:
 
$$\text{Zn}(s) + 2\text{HCl}(aq) \rightarrow \text{H}_2(g) + \text{ZnCl}_2(aq)$$

Which set of reaction conditions produces  $\text{H}_2(g)$  at the fastest rate?

  - a 1.0-g lump of Zn(s) in 50. mL of 0.5 M HCl(aq) at 20.°C
  - a 1.0-g lump of Zn(s) in 50. mL of 0.5 M HCl(aq) at 30.°C
  - 1.0 g of powdered Zn(s) in 50. mL of 1.0 M HCl(aq) at 20.°C
  - 1.0 g of powdered Zn(s) in 50. mL of 1.0 M HCl(aq) at 30.°C
- For a given reaction, adding a catalyst increases the rate of the reaction by
  - providing an alternate reaction pathway that has a higher activation energy
  - providing an alternate reaction pathway that has a lower activation energy
  - using the same reaction pathway and increasing the activation energy
  - using the same reaction pathway and decreasing the activation energy

- Given the balanced equation representing a reaction:
 
$$\text{CH}_4(g) + 2\text{O}_2(g) \rightarrow 2\text{H}_2\text{O}(g) + \text{CO}_2(g) + \text{heat}$$

Which statement is true about energy in this reaction?

  - The reaction is exothermic because it releases heat.
  - The reaction is exothermic because it absorbs heat.
  - The reaction is endothermic because it releases heat.
  - The reaction is endothermic because it absorbs heat.

- Which balanced equation represents an endothermic reaction?
  - $\text{C}(s) + \text{O}_2(g) \rightarrow \text{CO}_2(g)$
  - $\text{CH}_4(g) + 2\text{O}_2(g) \rightarrow \text{CO}_2(g) + 2\text{H}_2\text{O}(l)$
  - $\text{N}_2(g) + 3\text{H}_2(g) \rightarrow 2\text{NH}_3(g)$
  - $\text{N}_2(g) + \text{O}_2(g) \rightarrow 2\text{NO}(g)$
- Given the balanced equation:
 
$$\text{KNO}_3(s) + 34.89 \text{ kJ} \xrightarrow{\text{H}_2\text{O}} \text{K}^+(aq) + \text{NO}_3^-(aq)$$

Which statement best describes this process?

- It is endothermic and entropy increases.
  - It is endothermic and entropy decreases.
  - It is exothermic and entropy increases.
  - It is exothermic and entropy decreases.
- Which change is exothermic?
    - freezing of water
    - melting of iron
    - vaporization of ethanol
    - sublimation of iodine
  - Which reaction releases the greatest amount of energy per 2 moles of product?
    - $2\text{CO}(g) + \text{O}_2(g) \rightarrow 2\text{CO}_2(g)$
    - $4\text{Al}(s) + 3\text{O}_2(g) \rightarrow 2\text{Al}_2\text{O}_3(s)$
    - $2\text{H}_2(g) + \text{O}_2(g) \rightarrow 2\text{H}_2\text{O}(g)$
    - $\text{N}_2(g) + 3\text{H}_2(g) \rightarrow 2\text{NH}_3(g)$
  - Given the balanced equation representing a reaction at 101.3 kPa and 298 K:
 
$$\text{N}_2(g) + 3\text{H}_2(g) \rightarrow 2\text{NH}_3(g) + 91.8 \text{ kJ}$$

Which statement is true about this reaction?

    - It is exothermic and  $\Delta H$  equals  $-91.8 \text{ kJ}$ .
    - It is exothermic and  $\Delta H$  equals  $+91.8 \text{ kJ}$ .
    - It is endothermic and  $\Delta H$  equals  $-91.8 \text{ kJ}$ .
    - It is endothermic and  $\Delta H$  equals  $+91.8 \text{ kJ}$ .
  - Which expression represents the  $\Delta H$  for a chemical reaction in terms of the potential energy, PE, of its products and reactants?
    - PE of products + PE of reactants
    - PE of products – PE of reactants
    - PE of products = PE of reactants
    - PE of products + PE of reactants

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**JA Banks**



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### **Table of Contents Kinetics And Equilibrium Regents Questions Answers**

1. Understanding the eBook Kinetics And Equilibrium Regents Questions Answers
  - The Rise of Digital Reading Kinetics And Equilibrium Regents Questions Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Kinetics And Equilibrium Regents Questions Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Kinetics And Equilibrium Regents Questions Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Kinetics And Equilibrium Regents Questions Answers

- Personalized Recommendations
  - Kinetics And Equilibrium Regents Questions Answers User Reviews and Ratings
  - Kinetics And Equilibrium Regents Questions Answers and Bestseller Lists
5. Accessing Kinetics And Equilibrium Regents Questions Answers Free and Paid eBooks
    - Kinetics And Equilibrium Regents Questions Answers Public Domain eBooks
    - Kinetics And Equilibrium Regents Questions Answers eBook Subscription Services
    - Kinetics And Equilibrium Regents Questions Answers Budget-Friendly Options
  6. Navigating Kinetics And Equilibrium Regents Questions Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Kinetics And Equilibrium Regents Questions Answers Compatibility with Devices
    - Kinetics And Equilibrium Regents Questions Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Kinetics And Equilibrium Regents Questions Answers
    - Highlighting and Note-Taking Kinetics And Equilibrium Regents Questions Answers
    - Interactive Elements Kinetics And Equilibrium Regents Questions Answers
  8. Staying Engaged with Kinetics And Equilibrium Regents Questions Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Kinetics And Equilibrium Regents Questions Answers
  9. Balancing eBooks and Physical Books Kinetics And Equilibrium Regents Questions Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Kinetics And Equilibrium Regents Questions Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Kinetics And Equilibrium Regents Questions Answers
    - Setting Reading Goals Kinetics And Equilibrium Regents Questions Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Kinetics And Equilibrium Regents Questions Answers

- Fact-Checking eBook Content of Kinetics And Equilibrium Regents Questions Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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