

SOLUTIONS

The Laws of Thermodynamics

- **0th Law**
Introduces the thermodynamic intensive variable of temperature (T)
- **1st Law**
Conservation and conversion of energy
Defines extensive thermodynamic state variable of internal energy (U)
 $dU = dq + dw$
- **2nd Law**
Defines the extensive thermodynamic state variable of entropy (S)
 $dS_{univ} \geq 0$
- **3rd Law**
For systems in internal equilibrium, sets the zero of entropy at the minimum in temperature (0K) and at the minimum in internal energy

The Three Ed Equations

$$dS = \frac{dq_{rev}}{T} = \frac{dU + PdV - \sum \mu_i dn_i}{T}$$

$$dU = TdS - PdV + \sum \mu_i dn_i$$

$$dG = -SdT + VdP + \sum \mu_i dn_i$$

$$dA = -SdT - PdV + \sum \mu_i dn_i$$

$$dH = TdS + VdP + \sum \mu_i dn_i$$

$$dU = TdS - PdV + \sum \mu_i dn_i$$

$$dH = TdS + VdP + \sum \mu_i dn_i$$

$$dG = -SdT + VdP + \sum \mu_i dn_i$$

$$dA = -SdT - PdV + \sum \mu_i dn_i$$

Maxwell Relations for Single Component

$$\left(\frac{\partial T}{\partial V}\right)_S = -\left(\frac{\partial P}{\partial S}\right)_U$$

$$\left(\frac{\partial T}{\partial P}\right)_S = \left(\frac{\partial V}{\partial S}\right)_U$$

$$\left(\frac{\partial S}{\partial V}\right)_T = \left(\frac{\partial P}{\partial T}\right)_U$$

$$\left(\frac{\partial S}{\partial P}\right)_T = -\left(\frac{\partial V}{\partial T}\right)_U$$

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Minjie Lin



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