

Exercises, Questions, Model try – Preparation (Pre-Exam)

NOTE: (only) an unabbreviated nomenclature, a name) or (abbreviated nomenclature, IUPAC) is sufficient (unless otherwise stated, IUPAC nomenclature)

1. BaCl_2 and NaCl are dissolved in 100 mL of water, added to this 10% concentration (w/v) of the solution?
2. How many g of sulfate are in 100 g of solution if the concentration is 10% (w/w)?
3. A solution is composed with 10% (v/v), how many mL of the solution do you need in 100 mL of the solution?
4. What is the molarity (M) of potassium-permanganate by dissolving 0.1 g of KMnO_4 in 100 mL of water?
5. What is the molarity (M) of potassium-permanganate by dissolving 0.1 g of KMnO_4 in 0.1 L of water? (Molar mass $\text{KMnO}_4 = 158.034\text{ g/mol}$)
6. If a solution has a molarity of 1.0 M , how many L of it do you need to have 0.2 mol of the solute?
7. How many grams of MgCl_2 will you need to dissolve in 100 mL of water to create a solution with a molarity of 0.1 M ? (Molar mass $\text{MgCl}_2 = 95.21\text{ g/mol}$)
8. How many grams of sodium-sulfate will you need to create a solution with a molarity of 0.1 M ? How many mL of the stock solution (with 10% (w/v) of sodium SO_4) will you need to create?

Solutions Molarity And Dilution Practice Answer Key

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