

## Chapter 6 Chemical Bonds

**Section 6.4 The Structure of Metals****(pages 176–181)**

*This section discusses metallic bonds and the properties of metals. It also explains how the properties of an alloy are controlled.*

**Reading Strategy (page 176)**

**Relating Cause and Effect** As you read, complete the concept map to relate the structure of metals to their properties. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.



- Circle the letter of the metal with the highest melting point.
  - gold
  - vanadium
  - titanium
  - tungsten
- Is the following sentence true or false? The properties of a metal are related to bonds within the metal. true

**Metallic Bonds (pages 176–177)**

- Describe a metallic bond. A metallic bond is the attraction between a metal cation and the cation and the shared electrons that surround it.
- The cations in a metal form a lattice. What holds the lattice in place? The lattice is held in place by strong metallic bonds between the cations and the surrounding valence electrons.
- Is the following sentence true or false? The more valence electrons a metal has, the stronger its metallic bonds will be. true

**Explaining Properties of Metals (page 177)**

- Some of the properties of metals can be explained by the mobility of the electrons within a metal lattice.
- Name two important properties of metals that can be explained by metallic bonding.
  - Malleability
  - Ability to conduct an electric current

**Alloys (pages 178–181)**

- Circle the letter of the percentage of gold in jewelry that is labeled 18-karat gold.
  - 18 percent
  - 50 percent
  - 75 percent
  - 100 percent

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