

Section 6.1: Organizing the Elements

- Group 1 on the periodic table is called **Alkali Metals**
- Group 2 on the periodic table is called **Alkaline Earth Metals**
- Elements in groups 3-12 are called **Transition Metals**
- Group 17 on the periodic table is called **Halogens**
- Group 18 on the periodic table is called **Noble Gases**
- Each period on the periodic table corresponds to a **principal energy level**
- The 2 series that make up the inner transition metals are **Lanthanides** and **Actinides**
- Complete the following chart:

| Element | Symbol | Group # | Period # | Group Name | Metal, Metalloid, Nonmetal |
|----------------|-----------|-----------|----------|------------------------------|----------------------------|
| Zirconium | Zr | 4 | 5 | Transition Metals | Metal |
| Radium | Ra | 2 | 7 | Alkaline Earth Metals | Metal |
| Bismuth | Bi | 15 | 6 | Nitrogen Group | Metal |
| Gallium | Ga | 13 | 4 | Boron Group | Metal |

- What are triads? **Sets of 3 elements with similar properties** Whose idea were they? **Dobereiner**
- The modern periodic table is arranged by increasing **atomic number**. Whose idea was this? **Moseley**
- Mendeleev** developed the first periodic table, which was arranged by increasing **atomic mass**
- Determine which of the following pairs will have similar properties (Yes or no):
 - Ca and Be? **Yes – same group (Group 2 – Alkaline Earth Metals)**
 - Sr and Y? **No – not the same group**
 - C and Ge? **Yes – same group (Group 14 – Carbon Group)**

Section 6.3: Periodic Trends

- As you go across a period, the atomic radius (size) **decreases** because **the attraction between protons and electrons increases, which makes the atom smaller (opposite of tug of war)**
- As you go down a group, the atomic radius (size) **increases** because **the number of principal energy levels increases**
- Circle the atom with the largest atomic radius and put a square around the atom with the smallest radius:
 - | | | | |
|----|----|----|----|
| Li | Ne | F | B |
| N | Bi | S | P |
| Na | Sr | Al | Cl |
- Ionization energy is the energy required to **remove** an electron.

Modern Chemistry Chapter 6 Section 5 Review Answers

Katrin Zwirglmaier



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